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(54) Title: PHYTATE-ANTIMICROBIAL COMPOSITIONS IN ORAL CARE PRODUCTS (57) Abstract Oral compositions containing phytic acid or a physiologically acceptable salt thereof, a cationic antimicrobial compound and a compatibilizing agent are disclosed for controlling dental calculus, dental plaque, gingivitis, periodontitis and/or oral mal-odor.		

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PHYTATE-ANTIMICROBIAL COMPOSITIONS
IN ORAL CARE PRODUCTS

BACKGROUND OF THE INVENTION

5 The present invention relates to oral compositions containing an anticalculus or an antiplaque/antigingivitis agent.

 "Oral composition" means a composition for topical applications to the oral cavity to clean and care for the teeth as well as the oral cavity surfaces. Representatives of such compositions are oral hygiene products and dentifrices such as mouthwashes or rinses, toothpaste, dental gels, tooth powder, chewing gum, lozenges, and similar products. In addition to cleaning teeth to remove dental plaque, the function of oral hygiene preparations is to stop the formation of dental calculus, to prevent dental disorders such as caries, periodontitis and gingivitis, and also to eliminate halitosis.

 Dental calculus, or tartar as it is sometimes called, is a hard mineralized material which forms on teeth that consists of inorganic and organic components. The inorganic portion is largely calcium and orthophosphate arranged in a crystal lattice called hydroxyapatite (HAP). The organic portion is derived mainly from microorganisms (i.e., bacteria, yeast, etc.) as well as epithelial cells, white blood cells and food debris.

 Formation of dental calculus occurs in two steps. In the first step, plaque is deposited on the teeth. "Plaque" consists of inorganic and organic components derived from saliva, food and bacteria which are present in the oral cavity. Most of the plaque consists of dead and living bacteria surrounded by a gel-like matrix derived from the bacteria and saliva. In the second phase, plaque undergoes calcification to form dental calculus. Initially, amorphous deposits of calcium phosphate begins to appear on and within the matrix of the dental plaque. As the aggregates of calcium phosphate become sufficiently closely packed together, they crystallize to form HAP. The amorphous calcium phosphate, although related to hydroxyapatite, differs from it in crystal structure, particle morphology and stoichiometry.

 In addition to being an integral step for the formation of calculus, consequences of the presence of plaque include gingivitis, periodontitis, tooth decay (dental caries) and

denture associated problems. Inhibition of oral bacteria involved in the formation of plaque by antimicrobials or antiseptic agents is one means to retard the formation of plaque, thus aiding in preventing or controlling the formation of calculus and other plaque related diseases; see, for example, P. S. Hull, *J. Clinical Periodontology*, 7, 431-442 (1980). Examples of antiseptic agents include bisbiguanides, such as chlorhexidine and alexidine, and numerous antibacterially active quaternary ammonium compounds, such as cetylpyridinium chloride or the quaternary ammonium compounds described in U.S. Patents 3,369,046 and 4,820,507; and quaternary ammonium organosiloxane compounds described in U.S. Patent 4,161,518.

Although the quaternary ammonium compounds are rapidly adsorbed to the tooth surfaces, they exhibit only a moderate degree of efficacy as antiplaque and antigingivitis agents as they are rapidly released from the tooth surface and thus retained in the oral cavity for only a short period of time. Chlorhexidine has been the most successful antiplaque agent as it is believed to bind to the oral mucosa and is thus retained in the oral cavity for a longer period of time than quaternary ammonium compounds. The use of chlorhexidine in oral preparations however, suffers from the following disadvantages: (1) a prolonged bitter after taste lasting up to several hours; (2) after prolonged use stains are produced on the teeth, tongue, gums, oral mucosa and dental restorations; and (3) production of local irritation of the oral mucosa and tongue.

Inhibition of crystalline HAP formation is usually achieved by compounds which sorb onto a growing crystal and disrupt crystal growth. It is well known in the prior art that water soluble hexametaphosphates, tripolyphosphates and pyrophosphates and the like, are effective calcium and magnesium ion sequestrants and/or chelating agents. See, for example, U.S. Patent 3,488,419 which discloses oral compositions containing polyphosphate and U.S. Patent 4,215,105 which discloses oral compositions containing phosphonoacetic acid. However, as described in U.S. Patent 4,627,977, the effectiveness of polyphosphates as anticalculus agents has been limited because they are significantly hydrolyzed by salivary enzymes (phosphatases) to orthophosphates which are ineffective as inhibitors of HAP formation. The amount of enzymatic hydrolysis of the polyphosphate has been reduced by the use of a linear molecularly dehydrated polyphosphate salt combined with fluoride as described in U.S. Patent 4,808,410.

Compounds containing a carbon atom covalently bonded to oxygen, the oxygen being covalently bonded to a phosphorous, herein referred to as C-O-P bonds, particularly six C-O-P bonds, such as phytic acid [myo-inositol 1,2,3,4,5,6-hexakis(dihydrogen phosphate)], have been recommended for various purposes in oral compositions. U.S. Patents 4,259,316 and 4,335,102 disclose oral anticaries compositions containing a phytate compound and a stannous compound. Due to complex formation between polyvalent cations and phytate anion, the art teaches the presence of stannous compounds in an oral composition containing a phytate compound would not be desirable for inhibition of calculus formation.

In U.S. Patent 3,934,002 phytic acid is disclosed as one of the anticalculus compounds in oral compositions used together with a bisbiguanide antiplaque and anticaries agent. These two agents react with one another so that neither the anticalculus or antiplaque agent would be homogeneously distributed throughout the oral compositions. Since both agents are present, if a mouthrinse is prepared, it contains two visibly distinct phases, one being solid phase reaction product of bis-biguanide and anticalculus agent. U.S. Patents 4,263,276 and 4,305,928 also describe visually clear oral compositions containing phytic acid in the presence of an alkali metal fluoride, monofluorophosphate or alkali metal monofluorophosphate, where a cationic material, such as a bisbiguanide or cationic surface active agent can be present. However, effective inhibition of HAP or plaque formation with compositions including compounds containing C-O-P bonds with an antimicrobial have not been known beyond such recommendations or speculation.

It would therefore be desirable to have an oral composition containing an effective antiplaque or anticalculus agent to aid in the prevention of dental caries and gingivitis as well as aid in the control of mouth malodor which does not stain the teeth and does not have a bitter taste. It would also be desirable to provide an antiplaque and anticalculus oral composition in which phytic acid is homogeneously distributed in the oral composition along with a cationic antimicrobial compound.

Furthermore, it would be desirable to provide an oral composition having enhanced retention of cationic antimicrobial compounds on the tooth surfaces useful in the prevention of dental plaque and gingivitis.

SUMMARY OF THE INVENTION

The present invention relates to oral compositions containing phytic acid and a cationic antimicrobial compound and their use for the prevention of dental plaque. In particular, the present invention relates to an oral composition comprising: an orally acceptable vehicle containing: (a) from about 0.001 to about 10 percent by weight of one or more compounds having C-O-P bonds wherein the compound having C-O-P bonds is myo-inositol hexakis(dihydrogen phosphate), myo-inositol pentakis(dihydrogen phosphate), myo-inositol tetrakis(dihydrogen phosphate) or physiologically acceptable salts thereof; (b) from about 0.001 to about 10 percent by weight of one or more cationic antimicrobial compounds; (c) from about 0.1 to about 20 percent by weight of one or more compatibilizing agents; and (d) the remaining percent by weight is an orally acceptable vehicle.

A further embodiment of the present invention provides an improved method of inhibiting the formation of dental plaque and/or gingivitis and/or periodontitis and/or aid in controlling oral malodor.

A further embodiment of the present invention provides an oral composition containing phytic acid and a cationic antimicrobial compound which remain in solution, even in the presence of a polyvalent cation or polyvalent cations.

DETAILED DESCRIPTION OF THE INVENTION

5 The present invention relates to an oral composition consisting essentially of an orally acceptable vehicle containing phytic acid or derivatives thereof, a cationic component and a sufficient amount of a compatibilizing agent to inhibit the phytic acid and antimicrobial compound from interacting to form a precipitate when together in an aqueous solution. An "orally acceptable vehicle" means a medium in which an anticalculus or antiplaque agent may
10 be administered to the oral cavity surfaces without substantial harmful effects to the surfaces thereof.

As plaque is a main etiological factor in gingivitis, periodontitis, tooth decay (dental caries) and other dental associated problems, the ability to control dental plaque aids in preventing and/or controlling gingivitis, periodontitis and dental caries. Thus, as used herein,
15 "antiplaque" means antiplaque and/or antigingivitis and/or antiperiodontitis and/or anticaries. In addition, as the volatile sulfur compounds associated with oral malodor are related to the gingival health, as well as being produced by the putrefactive activity of microorganisms, as used herein, an antiplaque agent will also aid in the control of oral malodor.

To enhance the effect of preventing the formation of dental plaque, it has now
20 been unexpectedly found that the retention of cationic antimicrobial compounds and physiologically acceptable salts thereof to a tooth surface can be substantially enhanced if the cationic antimicrobial compound is used in combination with phytic acid or a derivative thereof in the presence of a sufficient amount of a compatibilizing agent to prevent the phytic acid and cationic antimicrobial compound from interacting to form a precipitate when exposed to each
25 other in an aqueous environment. The ability of an antiplaque agent to remain in contact with the tooth surface to exert an anti-plaque effect is referred to as "substantivity" of the agent. It has also been unexpectedly found that in the presence of a compatibilizing agent, the phytic acid and the cationic antimicrobial compound will remain in solution in the presence of polyvalent cations provided the ratio of polyvalent cation to phytic acid is not greater than
30 about 5 to 1.

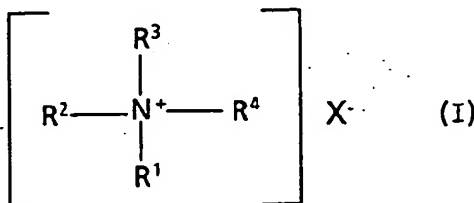
In certain preferred forms of the invention, the composition is substantially liquid in character, such as a mouthwash or rinse. In such a preparation the vehicle can be water or a water-alcohol mixture. When using a water-alcohol mixture, the weight ratio of water to alcohol is in the range of from about 1:1 to about 20:1, preferably about 3:1 to about 10:1 and
35 more preferably about 4:1 to about 6:1. The total amount of water or water-alcohol mixture in this type of preparation is typically in the range of from about 70 to about 99.9 percent by weight of the preparation. The pH of such liquid, and other liquid preparations of the

invention, is generally in the range of from about 4.5 to about 9, and typically from about 5.5 to about 8. The pH is preferably in the range of from about 6 to about 8.

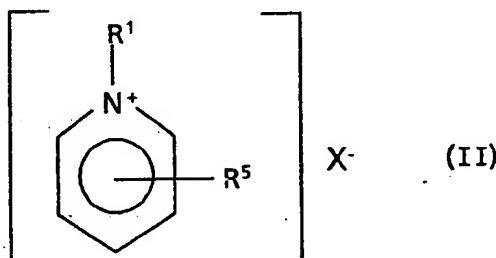
In certain other desirable forms of this invention, the oral composition may be substantially solid or semisolid in character, such as toothpowder, a dental tablet, a toothpaste, gel or dental cream. The vehicle of such solid or semisolid oral preparation generally contains added polishing material more fully described hereinafter.

As used herein, a "cationic antimicrobial compound" refers to an organic amine where the nitrogen is capable of being positively charged in an aqueous environment, and is represented by one or more of the following general formulae of A-J:

(A) Quaternary ammonium compounds represented by Formula I



or Formula II



wherein:

R¹ is a C₈-C₂₀ alkyl;

R² is benzyl or C₁-C₁₂ alkyl;

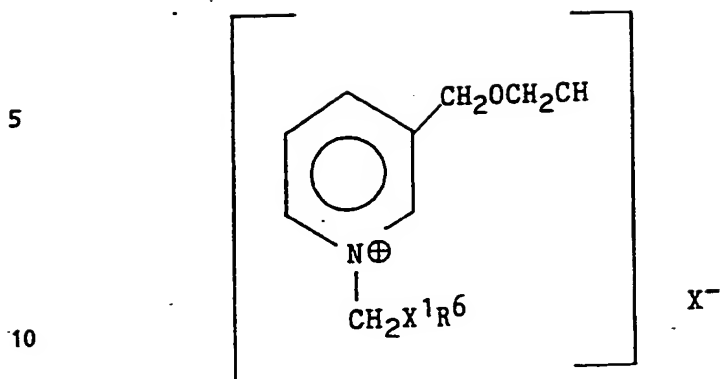
R³ and R⁴ are independently a C₁-C₇ alkyl or -(CH₂-CHOH-CH₂-O)_nH wherein n is an integer from 1 to 6;

R⁵ is -H, a C₁-C₇ alkyl or -(CH₂-CHOH-CH₂-O)_nH wherein n is an integer from 1 to 6;

and

X⁻ is chloride (Cl⁻), bromide (Br⁻), iodide (I⁻) or fluoride (F⁻) ion;

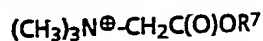
(B) Pyridinium chlorides containing alkylthiomethyl or alkoxymethyl hydrophobic groups as disclosed by Weglowski et al., *J. Phar. Sci.*, 80: 91-85 (1991), the disclosure of which is hereby incorporated by reference, having the formula



wherein X is as defined herein before and X¹ is oxygen or sulfur; and

R⁶ is a C₄-C₁₆ alkyl or benzyl;

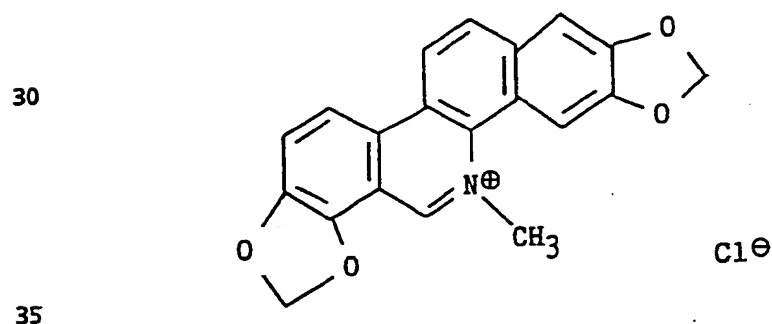
15 (C) Quaternary ammonium compounds that are esters of betaine and fatty alcohols, as disclosed by Linstead et al., *Antimicrobial Agents and Chemotherapy*, 39, 1949-1954 (1990), the disclosure of which is hereby incorporated by reference, having the formula



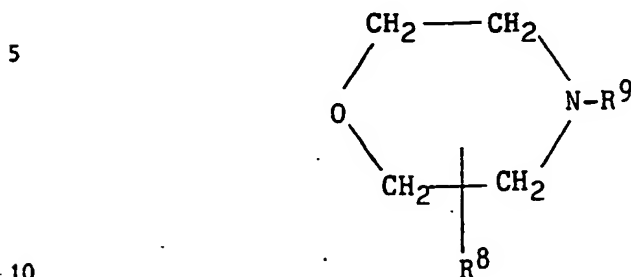
20 wherein R⁷ is a C₁₀-C₁₈ alkyl; and physiologically acceptable salts thereof;

(D) Sanguinarine and sanguinaria, sanguinaria being an extract from the bloodroot plant *Sanguinaria canadensis*, the extract containing benzophenanthridine alkaloids such as sanguinarine, chelerythrine, protopine, homochelidonine and physiologically acceptable salts thereof as disclosed in U.S. Patents 4,145,412 and 4,406,881, the disclosures of which are hereby incorporated by reference, sanguinaria being available in dentifrices under

25 the trademark Viadent™ brand sanguinaria; the major active ingredient sanguinarine chloride salt having the formula



(E) Morpholine compounds as disclosed in U.S. Patent 4,894,221, the disclosure of which is hereby incorporated by reference, the morpholine compounds having the formula

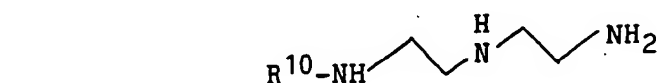


wherein R⁸ is a C₈-C₁₆ alkyl at the 2 or 3 position of the morpholino ring;

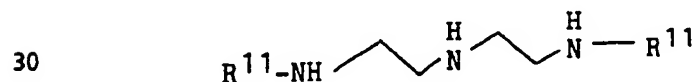
R⁹ is a C₂-C₁₀ alkyl substituted with a hydroxy group at other than the
15 alpha-position;

the sum of R₈ and R₉ being greater than or equal to 10 and preferably 10-20; and
physiologically acceptable salts thereof;

(F) Antibacterial secondary amines and amides as disclosed in *J. Antibacterial and Antifungal Agents*, 17, 371 (1989), the disclosure of which is hereby incorporated by reference,
20 wherein the antibacterial compounds have the following formulae

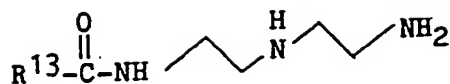


wherein R¹⁰ is a C₁₀-C₁₈ alkyl;



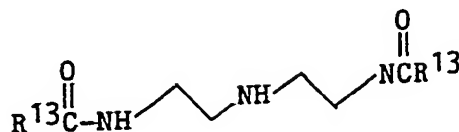
wherein each R¹¹ is independently C₈H₁₇ or C₁₀H₂₁;

35



wherein R¹³ is a C₉-C₁₇ alkyl;

or

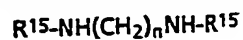


wherein each R¹³ is independently C₇H₁₅ or C₉H₁₉; and physiologically acceptable salts thereof;

(G) Dialkyl amines and N,N'-dialkylpolymethylene-diamines as disclosed in *J. Antibacterial and Antifungal Agents*, 17, 579 (1989), the disclosure of which is hereby incorporated by reference, having the formula



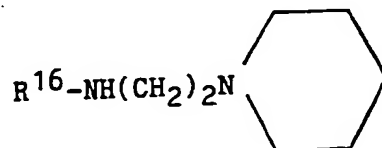
wherein each R¹⁴ is independently C₈H₁₇ or C₁₂H₂₅; or formula



wherein each R¹⁵ is independently a C₇-C₁₀ alkyl;

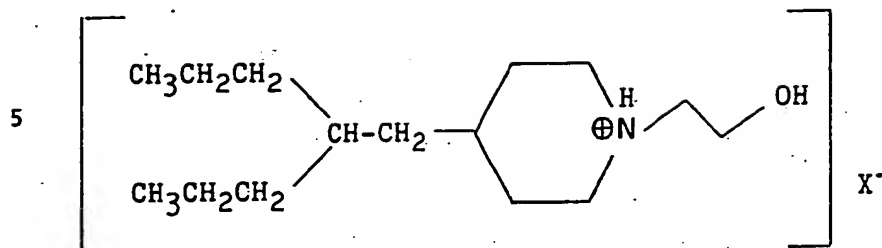
n is an integer from 2 to 5; and physiologically acceptable salts thereof;

(H) N'-Alkyl-N-(2-aminoethyl)piperidine compounds as disclosed by Murata et al., *J. Pharm. Sci.*, 80, 26-28 (1991), the disclosure of which is hereby incorporated by reference, the compounds having the formula



wherein R¹⁶ is a C₁₀-C₁₈ alkyl; and physiologically acceptable salts thereof;

(I) The ammonium compound 4-(2-propylenepentyl)-1-piperidinoethanol described in *J. Periodontal Research*, 18, pp. 429-437 (1983), the compound having the structure

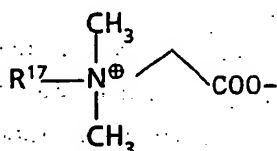


10

wherein X^- is as defined hereinbefore; described in the literature as Octapinal™ brand 4-(2-propylenepentyl)-1-piperidinoethanol (Ferrosan AB, Sweden); and

(J) Alkyl-N-betaine in combination with an alkyl-N,N-dimethylamine oxide; the alkyl-N-betaine having the structure

15

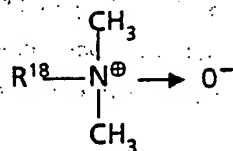


20

wherein R^{17} is a C_{10} - C_{18} alkyl;

the alkyl-N,N-dimethylamine oxide having the structure

25



30

wherein R^{18} is a C_{10} - C_{18} alkyl;

as disclosed in U.S. Patent 4,839,158, the disclosure of which is hereby incorporated by reference.

35

As used herein, the term "alkyl" means a linear or branched alkyl and thus secondary and tertiary alkyls are included. The alkyl terms up to C_{20} include, for example, t-butyl, sec-butyl, isobutyl, and in like manner all such branched or straight chain alkyls.

Preferred quaternary ammonium antibacterial agents include dodecyl trimethyl ammonium bromide, benzyl dimethyl stearyl ammonium chloride, N-tetradecyl-

-4-ethylpyridinium chloride and cetylpyridinium chloride. The terms antibacterial and antimicrobial mean the ability to inhibit growth; metabolism or reproduction of microorganisms.

The cationic antimicrobial compounds useful in the present invention are commercially available or may be obtained by those of ordinary skill in the art without undue experimentation. For example, quaternary ammonium compounds may be produced by reacting alkyl halides with ammonia or primary amines, or by reacting a tertiary amine, pyridine or pyridine derivative with an alkyl halide. See, for example, Zoltewicz and Deady, *Adv. Heterocycl. Chem.*, 22, 71-121 (1978); U.S. Patents 2,446,792; 2,295,504 and 4,994,199, the teachings of which are hereby incorporated by reference.

One or more cationic antimicrobial compounds are employed in amounts such that the oral product contains from about 0.001 and 10 percent by weight of the antimicrobial compound. Preferably for desired levels of antiplaque and antigingivitis effect, the finished oral product contains about 0.01 to about 5 percent and preferably 0.025 to 1.0 percent by weight of the antimicrobial compound. Typically a singular antimicrobial compound is employed in the oral product.

The compounds of the present invention which contain C-O-P bonds are phosphate esters of myo-inositol, such as phytic acid, also known as myo-inositol hexakis(dihydrogen phosphate), inositol hexaphosphoric acid, and 1,2,3,4,5,6-cyclohexanehexaol-phosphoric acid. As used herein "phytic acid" means the hexakis phosphate ester of myo-inositol, myo-inositol hexakis(dihydrogen phosphate), and the lesser substituted tetrakis and pentakis phosphate esters of myo-inositol, myo-inositol tetrakis(dihydrogen phosphate) and myo-inositol pentakis (dihydrogen phosphate) respectively, and physiologically acceptable salts thereof, such as alkali metal, alkaline-earth metal, ammonium salts or mixtures thereof. These phytic acid compounds may be used singly or in combination. Phytin, which is the calcium magnesium salt of phytic acid represented by the formula $\text{Ca}_5\text{Mg}(\text{C}_6\text{H}_{12}\text{O}_{24}\text{P}_6 \cdot 3\text{H}_2\text{O})_2$, can also be used in the present invention in addition to or replacement of the phytic acid.

Phytic acid and phytin are commercially available. The tetrakis and pentakis phosphate esters of inositol compounds can be prepared by hydrolyzing sodium phytate with hydrochloric acid and separating the inositol phosphates by high performance liquid chromatography as described by Sanberg and Ahderinne, *J. Food Sci.*, 51, 547-550 (1986), the disclosure of which is hereby incorporated by reference.

Phytic acid is present in the oral composition of the present invention, in an amount from about 0.001 to about 10 percent by weight. When the oral composition is essentially liquid in nature, the phytic acid or salt is typically present in an amount from about 0.001 to about 10 percent by weight, preferably from about 0.005 to about 5 percent and more preferably from about 0.01 to about 1 percent by weight.

When the oral composition is essentially liquid in nature, to maintain the cationic antimicrobial compound and phytic acid in solution, it is desirable that the composition contain a sufficient amount of a compatibilizing agent to keep the phytic acid and cationic antimicrobial compound from interacting to form a precipitate. Compatibilizing agents in the present invention are those which do not have a detrimental effect on the substantivity of the cationic antimicrobial compound in the presence of phytic acid and maintains the cationic antimicrobial compound and phytic acid in solution when the oral composition is essentially liquid in nature such that an aqueous solution of the phytic acid, cationic antimicrobial compound and compatibilizing agent does not visually become turbid after standing for 30 minutes at room temperature. A detrimental effect on substantivity means the retention of the antimicrobial compound near the tooth surface in the presence of phytic acid is not substantially different than the retention of the antimicrobial compound in the absence of phytic acid. It is therefore possible that a compound or combination of compounds may combatilize the phytic acid and antimicrobial compound, i.e., keep them in solution, but adversely affect the substantivity of the antimicrobial/phytic acid solution.

When the oral composition is substantially gel-like or semisolid in form, the vehicle of such solid oral preparation contains a liquid moiety of water so that the phytic acid, cationic antimicrobial compound and compatibilizing agent are homogeneously distributed throughout the liquid phase of the composition. The total amount of water in gel-like or semisolid-like oral compositions is typically in the range of from about 5 to about 60 percent by weight of the preparation; preferably from about 10 to about 50 percent.

While not wishing to be bound by theory, it is believed the compatibilizing agents of the present invention reduce the interaction between the phytic acid and cationic antimicrobial compound, reducing or preventing the formation of a precipitate when these two compounds are exposed to each other in an aqueous environment. The amount of compatibilizing agent in the oral compositions of the present invention is from about 0.1 to about 20 percent by weight, preferably from about 0.1 to about 10 percent by weight of the total composition. Particularly useful compatibilizing agents for oral compositions of the present invention which are substantially liquid in nature are acids and their alkali metal or alkaline-earth metal salts, or mixtures thereof. The mixtures are designated herein as anionic buffers. Suitable anionic buffers are, for example, phosphate, acetate, borate, citrate, bicarbonate, gluconate, tartrate, sulfate, and the like, or mixtures thereof. The preferred anionic buffers being phosphate and bicarbonate. When the oral composition is essentially in the liquid form, the anionic buffer is present in a concentration of from about 0.1 M to about 1.0 M, preferably from about 0.25 M to about 0.75 M.

Other examples of compatibilizing agents useful in the present invention are surfactants which maintain the phytic acid and antimicrobial compound in solution and does not interfere with enhanced substantivity due to the presence of phytic acid. An example of a

particularly suitable nonionic surfactant is poly(oxyethylene), poly(oxypropylene) block polymers known as poloxamers and available, for example, under the trademark "PLURONICS" (BASF Wyandotte Co., Parsippany, NJ). Another example of a particularly suitable nonionic surfactant is polyethylene oxide sorbitan esters, available for example, under the trademark "TWEENS" (ICI American Inc., Wilmington, De.) Suitable anionic surfactants include, for example, anionic surfactants produced from fatty acids and the amino acid sarcosine, such as N-lauroyl sarcosine, available for example, under the trademark "HAMPOSYL" by W.R. Grace and Co. (CT).

When utilizing a combination of compatibilizing agents, it is desirable that the total amount of compatibilizing agent in the oral composition remain from about 0.1 to about 20 weight percent. The concentration of compatibilizing agent or agents for use in the present invention can be readily determined by those of ordinary skill in the art based upon the teachings herein.

In a further embodiment, it has been found that when a polyethylene oxide sorbitan ester is used as a compatibilizing agent for phytic acid and an antimicrobial compound such as N-tetradecyl-4-ethylpyridinium bromide, the concentration of compatibilizing agent can be reduced below 0.1 weight percent, and about 0.05 weight percent can be used.

The molar ratio of phytic acid to the cationic antimicrobial compound in the presence of a compatibilizing agent is preferably from about 10:1 to about 1:10, more preferably from about 5:1 to about 1:5, and most preferably about 1:1.

Preparation of the oral compositions of the present invention can be made by using customary procedures for unifying components applied to the teeth and gingiva. It has been found that liquid mouthwashes and topical solutions of the present invention, can be prepared by: (a) dissolving the phytic acid and compatibilizing agent in water, (b) adjusting the pH to between about 6 to about 8, and then (c) adding the cationic antimicrobial compound. When the oral compositions of the present invention contain a polyvalent metal ion in addition to the phytic acid, the compatibilizing agent and the cationic antimicrobial compound, then the compositions are advantageously prepared by: (a) dissolving the metal ion, phytic acid, and a compatibilizing agent in water, (b) adjusting the pH to between about 6 to about 8, and then (c) adding the cationic antimicrobial compound. Alternatively, the cationic antimicrobial compound can be added to the solutions above, prior to adjusting the pH. Other components, such as sweetening and flavoring agents as described more fully herein, can then be added if desired.

To prepare an oral composition which is substantially solid or semisolid in character, an aqueous solution of the phytic acid, cationic antimicrobial compound and compatibilizing agent is prepared in the ratios as described above, and the water removed. Alternatively, a substantially solid or semisolid oral composition containing phytic acid, cationic antimicrobial compound and compatibilizing agent may be prepared by mixing the

components in the preferred ratios with the other ingredients of the oral composition as described herein.

It has been unexpectedly found that a metal ion selected from strontium (Sr^{2+}), magnesium (Mg^{2+}), tin (Sn^{2+}), zinc (Zn^{2+}), calcium (Ca^{2+}) or mixtures thereof, can be added to the oral composition containing the phytic acid, cationic antimicrobial compound and compatibilizing agent without the phytic acid precipitating from solution. The molar ratio of the metal ion to the phytic acid which can be present in the oral compositions of the present invention is from about 4:1 to about 1:4, preferably from about 3:1 to about 1:3, and more preferably about 1:1. The inclusion of a metal ion with the phytic acid and antimicrobial agent would aid in the suppression of oral malodor in addition to aiding in the control of calculus, plaque and gingivitis due to the inhibitory effect of the antimicrobial compound.

The dentifrices of the present invention may also be in a kit form for treating the oral cavity, the kit comprising phytic acid in an orally acceptable vehicle, one or more compatibilizing agents in an orally acceptable vehicle, and one or more cationic antimicrobial compounds in an orally acceptable vehicle; and a means to contain the phytic acid separately from the cationic antimicrobial. Means to separate the phytic acid and cationic antimicrobial includes placing them in separate vessels or in a compartmentalized container. The compatibilizing agent may be mixed with the phytic acid, with the cationic antimicrobial or may be contained separately.

When the dentifrice of the present invention is in a kit form, the phytic acid, compatibilizing agent and cationic antimicrobial compound is mixed prior to application.

When mixing the phytic acid, compatibilizing agent and cationic antimicrobial compound prior to application to the oral cavity, it may be necessary to increase their concentration to account for dilution effects which can occur upon mixing. When applying the phytic acid, compatibilizing agent and cationic antimicrobial in a kit form by mixing prior to use, the concentration of the individual compounds to which the oral cavity is exposed should be in the range given hereinbefore for their concentration in the final dentifrice product.

A variety of other ingredients may be added to the dentifrices of the present invention. Thus for example, prophylactic agents, polishing agents, soaps or detergents, flavoring and sweetening agents, thickening agents and humectants may be included using techniques which are known in the art.

Representative prophylactic agents include supplemental caries-preventing materials such as sodium fluoride, stannous fluoride, potassium fluoride, hexylamine hydrofluoride, myristylamine hydrofluoride, betaine fluoride, glycine potassium fluoride, etc. A particularly preferred fluoride is sodium fluoride. Typically these prophylactic agents are present in sufficient concentrations so as to provide an available fluoride ion concentration of up to about 2 percent by weight, and preferably in the range of from about 0.5 to about 2 percent by weight, of the dentifrice composition.

Suitable polishing agents include, for example, abrasive materials such as insoluble condensed phosphates such as calcium pyrophosphate, insoluble calcium polyphosphate (also known as calcium polymetaphosphate) and highly polymerized sodium polyphosphate; and water impervious cross-linked thermosetting resins such as the condensations products of melamine and urea with formaldehyde. Other suitable polishing agents will be obvious to those skilled in the art.

The polishing material is generally present in the solid or semisolid compositions in weight concentrations of from about 10 to about 99 percent. Preferably, it is present in amounts ranging from 20 to 75 percent in toothpaste, and from about 70 to about 99 percent in tooth powder.

Soaps or detergents may also be employed in the present invention to lower the surface tension to achieve increased prophylactic action, assist in achieving thorough and complete dispersion of the anticalculus agent and render the instant compositions more cosmetically acceptable. Suitable soaps include, for example, the soaps of high molecular weight fatty acids such as sodium and potassium soaps of myristic, stearic or palmitic acids and fatty acids mixtures of palm oil and coconut oil. Typical synthetic detergents include alkyl sulfates and sulfonates having alkyl groups of from about 8 to about 18 carbon atoms, such as, for example, sodium lauryl sulfate, the sulfated fatty alcohols derived from coconut oil and palm oil. The soaps typically comprise up to about 5 percent by weight of the dentifrice composition.

Any suitable flavoring or sweetening material may also be employed. Examples of suitable flavoring constituents are flavoring oils, e.g., oil of spearmint, peppermint, wintergreen, sassafras, clove, sage, eucalyptus, marjoram, cinnamon, lemon and orange and methyl salicylate. Suitable sweetening agents include sucrose, lactose, maltose, sorbitol, xylitol, sodium cyclamate, perillartine, APM (aspartylphenylalanine, methyl ester), saccharine and the like. Suitably, flavor and sweetening agents may together comprise from about 0.1 percent to 5 percent of the preparation.

Toothpastes, creams and gels typically contain a natural or synthetic thickener or gelling agent in proportions of from about 0.1 to about 10 percent, preferably from about 0.5 to about 5 percent, by weight. Suitable gelling or thickening agents include for example, water-soluble salts of cellulose ethers such as sodium carboxymethyl cellulose and sodium carboxymethylhydroxy-ethyl cellulose; natural gums such as gum karaya, gum arabic, and gum tragacanth; and colloidal magnesium aluminum silicate or finely divided silica.

Suitable humectants which may be employed in compositions of the invention include glycerine, propylene glycol, sorbitol, polypropylene glycol and/or polyethylene glycol and other polyhydric alcohols. The humectants may comprise from about 10 to about 90 percent by weight of the dentifrice composition.

The invention will be further clarified by a consideration of the following examples, which are intended to be purely exemplary of the present invention.

GENERAL EXPERIMENTAL

Preparation of Stock Solutions

5 0.045 M phytic acid. A 0.045 molar (M) solution of phytic acid was prepared by dissolving 1.462 g (1.125 millimoles (mMol)) of 50 percent by weight of phytic acid (Jonas Chemical Corp.) in 10 mL of water. The pH of this solution was brought to 7.78 by the addition of 1 N NaOH. This solution was transferred quantitatively to a 25 mL volumetric flask and diluted to the mark with water.

10 0.0045 M phytic acid. A 0.0045 M solution of phytic acid was prepared by adding 0.743 mL of phytic acid (40 percent weight solution, Aldrich Chemical Co., Inc.) to a 100 mL volumetric flask and diluting to the mark with water.

0.0045 M CPC. A 0.0045 M solution of cetylpyridinium chloride (CPC) was prepared by adding 0.4026 \pm 0.0001 g of CPC (Aldrich Chemical Co., Inc.) to a 250 mL volumetric flask, dissolving in

15 water and diluting to mark with water. The final pH of the solution was 7.4.
0.045 M HEDP. A 0.045 M solution of hydroxyethylidene-diphosphonic acid (HEDP) was prepared by adding 0.0618 \pm 0.0001 g of 60 percent active HEDP (MAYO Chemical Co.) to a 60 mL beaker and adding 40 mL of water. This solution was adjusted to pH 7.6 using a few drops of 1.0 N sodium hydroxide.

20 1.5 M phosphoric acid. A 1.5 M phosphoric acid solution was prepared by adding 25 mL of water to a beaker containing 17.29 g of 85 percent by weight solution of phosphoric acid (Mallinckrodt). The pH of this solution was raised to about 7 with the addition of 50 percent by weight sodium hydroxide. The solution was then transferred quantitatively to a 100 mL volumetric flask and diluted to the mark with water.

25 Glycolysis pH Test

A sucrose solution was prepared by loading 1.0 g of sucrose (Imperial Pure Cane Sugar) into a 60 mL beaker and then adding 20 mL of water. To this solution was added 8.0 mL of pooled whole human saliva. The saliva was collected from donors who had been permitted to eat or drink anything prior to collection period, but had foregone any oral hygiene on the

30 day of collection. Prior to the collection, each donor rinsed their mouth for thirty seconds with approximately 30 mL water, and after waiting about 5 minutes, began collecting saliva for 30 to 40 minutes, keeping the collected saliva on ice.

To the saliva/sucrose solution was added 1.0 mL of brain/heart infusion broth containing *Streptococcus mutants* (American Type Culture Collections No. 25175, ATCC) and

35 1.0 mL of brain/heart infusion broth containing *Streptococcus sanguis* (ATCC #10556). These cultures had been inoculated into 40 mL of broth and grown at 37°C for sixteen hours prior to adding to the saliva/sucrose solution. (Each broth contained approximately 60 million colony forming units at the time of addition.)

Aliquots of 0.75 mL of the above saliva/sucrose/bacterial solution were added to the test tubes containing various washed HAP suspensions. These test tubes were capped and attached to a tube rotator and placed in a 37°C incubator for sixteen hours. Following this incubation period, the rotator was removed from the oven and allowed to cool to ambient temperature. The pH of the solutions were checked with a pH meter using a pH electrode calibrated with pH 4, 7 and 10 buffers.

Treating and Washing Hydroxyapatite

To determine the substantivity of CPC in combination with phytic acid, the following washing procedure of the hydroxyapatite was done prior to performing a glycolysis pH test:

A 60 mL beaker was loaded with 12.0 g of hydroxyapatite (HAP) in a buffer suspension (25 percent by weight solids from Sigma Chemical Co.) and washed with 25 mL of water. The HAP suspension was filtered through a medium glass fitted filter to obtain a HAP filter cake. The HAP filter cake was washed a second time with an additional 25 mL of water and then filtered through a glass filter funnel. The white solid filter cake containing 3.0 g of HAP without the buffer was resuspended with 30.0 mL of water to produce a 3.0 g HAP/30.0 mL or 100 mg/mL suspension.

Two mL of the HAP suspension was transferred to each of several sterile-disposable polystyrene 5 mL test tubes labeled D_1-D_n (where n = number of test solutions). Two mL of a test solution were then added to each test tube.

The tubes containing the HAP and test solution were capped and attached to a tube rotator and rotated end-over-end to allow the test solutions to contact the HAP for a total of ten minutes.

After mixing, the test tubes were placed in an Industrial Equipment Company (IEC) model K centrifuge and spun at setting 25 (mid-range) for ten minutes. The tubes were removed and the liquid layer decanted. A macropipettor was then used to add 3.0 mL of water to each test tube containing the centrifuged hydroxyapatite. The HAP solids were resuspended by vigorous in-and-out flowing action through the pipette. The tubes were again centrifuged at setting 25 for ten minutes and the liquid layer decanted. Following the three milliliter wash step, the HAP solids were resuspended in 2.0 mL of water to produce the original 100 mg/mL suspension concentration. A 0.5 mL sample (containing 50 mg HAP) of this HAP suspension was removed and placed in each of several 5 mL polystyrene test tube labeled A_1-A_n . This sample A contains one fourth of the original HAP suspension which has been washed with three milliliters of water.

The remaining 1.5 mL in test tubes labeled D_1-D_n were centrifuged for ten minutes, the test tubes removed, and the liquid layer decanted. Three milliliters of water were added to these test tubes and the HAP solids resuspended/washed using disposable pipettes. The tubes were centrifuged for ten minutes, the tubes removed and the liquid layer decanted.

An additional three milliliters of water were added to these tubes and the HAP solids resuspended/washed by pipette. These tubes were again placed in the centrifuge and spun for ten minutes. The tubes were removed, the liquid layer decanted and 1.5 mL of water added to each tube. The HAP solids were resuspended to the original 100 mg/mL concentration and a 0.5 mL sample removed and placed in each of several 5 mL polystyrene test tubes labeled B₁–B_n. This sample B contained 50 mg HAP solids which had been treated with test solution and then washed with a total of eleven milliliters of water.

The procedure given above was repeated a third and fourth time to create a series of test tubes labeled C₁–C_n and D₁–D_n. The C samples contained HAP solids which had been treated with test solution and then washed with a total of 18.5 mL of water. The D samples contained HAP solids which had been treated with the test solution and then washed with a total of 25.5 mL water.

A glycolysis pH test was then performed as described above by adding 0.75 aliquots of the saliva/sucrose/bacterial mixture to the test tubes labeled A₁–D_n, each containing 0.5 mL of the treated washed HAP suspension.

Example 1: Incompatibility of CPC with phytic acid

Into each of eight vials was placed 500 µL of 0.0045 M CPC, varying amounts of a 0.045 M phytic acid solution and water as shown in Table I to give a final concentration of 1.5 mM CPC in all the samples. The phytic acid solution and water were combined first.

The vials were then capped and allowed to mix on an end-over-end rotator for 18 hours to allow for complete precipitation. At the end of this time, the tubes were centrifuged at 4,000 rpm for 5 minutes and a 200 µL aliquot of the supernatant removed and added to three mL of water. These diluted samples were then mixed in a quartz cuvette and the ultraviolet (UV) absorption determined at 260 nanometers. The amount of CPC still in solution was determined by comparing the absorbance of the samples to the absorbance of a 0.0015 M CPC solution. The results, as shown in Table I, show CPC precipitates in the presence of phytic acid.

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TABLE I

μL Phytic Acid	μL Water	mM CPC*	mM Phytic Acid	% of CPC* still in Solution
1000	0	1.5	30	21.3
500	500	1.5	15	10.7
100	900	1.5	3	3.5
50	950	1.5	1.5	1.9
33.3	966.7	1.5	1.0	1.8
26.7	973.7	1.5	0.8	1
10	990	1.5	0.3	2.9
5	995	1.5	0.15	13.4

*CPC = cetylpyridinium chloride

Example 2

The procedure in Example 1 was repeated except the CPC and phytic acid solutions were made in 0.5 M phosphate buffer at pH 7.3. The phosphate buffer was prepared from 85 percent by weight of phosphoric acid and adjusted to pH about 7.3 with 50 percent by weight of sodium hydroxide. The phosphate buffer was also substituted for the water portion of the solutions. The percent of CPC remaining in solution was determined by measuring the ultraviolet absorbance of the samples compared to a 0.0015 M CPC standard. The results, given in Table II, show no CPC is precipitated from solution at any of the various concentrations and demonstrate the compatibilization of CPC with phytic acid in the presence of an anionic buffer.

TABLE II -

μL Phytic Acid	μL Phosphate Buffer	mM Phytic Acid	% of CPC* still in Solution
1000	0	30	105.8
500	500	15	102.4
100	900	3	103.2
50	950	1.5	103.6
33.3	966.7	1.0	103
26.7	973.3	0.8	100.4
10	990	0.23	103.7
5	995	0.15	101.7

*CPC = cetylpyridinium chloride

Example 3

Using the experimental procedures described above for glycolysis test and preparation of the hydroxyapatite, the following compounds were tested for HAP

substantivity: water (control); cetylpyridinium chloride (CPC); phytin/phosphoric acid/citric acid (PyPCi); hydroxyethylidene phosphate (HEDP); hydroxyethylidene phosphate/cetylpyridinium chloride (HEDP/CPC); and phytin/phosphoric acid/citric acid/cetylpyridinium chloride (PyPCiC). All the compounds tested were 0.0015 M except for phosphate and citrate which were 0.015 M.

The 0.0015 M CPC and 0.0015 M HEDP solutions were prepared by diluting 0.0045 M stock solutions prepared as previously described.

The PyPC (phytin/phosphate/citrate) solution was prepared by placing 123 mg (75 μmoles) of phytin (American Tokyo Kasei, Inc.) and 4 mL of 0.15 M citric acid into a 4 ounce jar. (The citric acid being prepared by adding 1.575 g of citric acid to water and bringing the final volume to 50 mL in a volumetric flask). To this suspension of phytin and citric acid was added 2.873 g (0.025 mole) of neat phosphoric acid followed by 30 mL of water. The pH was adjusted to 7.38 by the addition of 50 percent by weight of sodium hydroxide. After allowing the solution to cool to room temperature, an additional 10 mL of water was added to bring the volume to 50 mL. The pH was 7.36. The solution had a faint turbidity and was centrifuged and the supernatant decanted giving a clear solution.

The PyPCIC (phytin/phosphate/citrate/cetyl-pyridinium chloride) solution was prepared by placing 5.37 mg of CPC into a 10 mL volumetric flask and bringing to the 10 mL mark using the phytin/phosphate/citrate solution described above.

5 The HEDP/CPC (hydroxyethylidene phosphate/cetylpyridinium chloride solution) was prepared by placing 20 mL of a 0.0045 M HEDP solution into a 4 ounce jar and adding 20 mL of water and 20 mL of a 0.0045 M CPC stock solution. The pH was adjusted to 7.4 with the dropwise addition of 1.0 N sodium hydroxide.

10 The phosphate buffer/CPC (Phos/CPC) solution was prepared by adding 20 mL of a 0.042 M sodium phosphate buffer solution to a four ounce jar and diluting with 20 mL of water and 20 mL of a 0.0045 M CPC stock solution. The 0.042 M sodium phosphate buffer was prepared by adding 0.2318 ± 0.0001 g of $\text{NaH}_2\text{PO}_4 \cdot \text{H}_2\text{O}$ (J.T. Baker Chemical Co.) to a 60 mL beaker and adding 40 mL of water and adjusting the pH to 7.36 by the dropwise addition of 1 N sodium hydroxide.

15 The results of the glycolysis pH test for HAP treated by the above compounds are shown Table III.

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TABLE III
pH as a Function of HAP Washings

Volume of Wash (mL)	Water	CPC ¹	PyPCi ²	HEDP ³	HEDP/ CPC ⁴	Phos/ CPC ⁵	PyPCiC ⁶
3	5.04	6.98	5.07	5.04	7.04	7.02	7.3
11	5.12	5.29	5.12	5.5	5.87	5.84	7.01
18.5	5.06	5.13	5.18	5.21	5.3	5.18	7.04
25.5	5.14	5.35	5.2	5.15	5.29	5.1	7.04

¹CPC = cetylpyridinium chloride

²PyPCi = phytin, phosphate and cetylpyridinium chloride

³HEDP = hydroxyethylidene phosphate

⁴HEDP/CPC = hydroxyethylidene phosphate and cetylpyridinium chloride

⁵Phos/CPC = phosphate and cetylpyridinium chloride

⁶PyPCiC = phytin, phosphate, citrate and cetylpyridinium chloride

These results show that retention of cetylpyridinium chloride as measured by the retention of antimicrobial activity was greatest for the PyPCiC solution.

Example 4

- Using the experimental procedures described above for the glycolysis test and preparation of the hydroxyapatite, the following compounds were tested for HAP substantivity: water (control); cetylpyridinium chloride (CPC); phytin/phosphoric acid/citric acid/cetylpyridinium chloride (PyPCiC); and phytic acid/phosphoric acid/citric acid/cetylpyridinium chloride (PaPCiC). The CPC and PyPCiC solutions were prepared as given in Example 3. The phytic acid/phosphoric acid/citric acid/cetylpyridinium chloride (PaPCiC) solution was prepared by adding 77.2 μ L (150 μ moles) of phytic acid (40 percent by weight, Aldrich Chemical Co., Inc.), 2.873 g of phosphoric acid, 4 mL of 0.15 M citric acid and 25 mL water to a vial. This solution was brought to pH 7.39 by the addition of 3.71 g of 50 percent by weight of sodium hydroxide and diluted to mark with water in a 50 mL volumetric flask. A 5.37 mg portion of CPC was dissolved in 10 mL of this solution to produce a solution containing 3 mM phytic acid, 12 mM citrate, 1.5 mM CPC and 500 mM phosphate.

The results of the glycolysis pH test for HAP treated by the above compounds are shown in Table IV.

TABLE IV
pH as a Function of HAP Washing

Volume of Wash (mL)	Water	CPC ¹	PyPCiC ²	PaPCiC ³
3	4.83	7.12	7.34	7.40
11	4.91	6.02	7.07	7.14
18.5	4.93	5.15	7.18	7.22
25.5	5.17	5.35	7.17	7.24

¹CPC = cetylpyridinium chloride

²PyPCiC = phytin, phosphate, citrate and cetylpyridinium chloride

³PaPCiC = phytic acid, phosphate, citrate and cetylpyridinium chloride

These results show that both phytin and phytic acid enhance the substantivity of CPC to hydroxyapatite as measured by the glycolysis pH test.

Example 5

Preparation of the HAP and the treatment of the HAP with test solutions was as described under general experimental. However, the amount of water wash of the HAP suspensions was increased from three to 10 mL. As a result, samples A, B, C and D were removed after washing with 10, 31, 51.5 and 71.5 mL of water, respectively. Using this wash procedure, the solutions as prepared in Example 4 were tested for HAP substantivity: water (control); cetylpyridinium chloride (CPC); phytin/phosphoric acid/citric acid/cetylpyridinium chloride (PyPCiC); phytic acid/phosphoric acid/citric acid/cetylpyridinium (PaPCiC).

The results of the glycolysis pH test for HAP treated by the above compounds are shown in Table V.

TABLE V
pH Drop as a Function of HAP Washings

mL of Water Wash	Water	CPC ¹	PyPCiC ²	PaPCiC ³
10	4.76	7.15	7.32	7.26
31	4.83	5.04	7.24	7.21
51.5	4.88	5.01	7.34	7.30
71.5	5.19	5.17	7.36	7.28

¹CPC = cetylpyridinium chloride

²PyPCiC = phytin, phosphate, citrate and cetylpyridinium chloride

³PaPCiC = phytic acid, phosphate, citrate and cetylpyridinium chloride

The results show that retention of cetylpyridinium chloride as measured by the retention of antimicrobial activity was substantially enhanced by the presence of phytic acid or phytin and an anionic buffer under extended washing procedures.

Example 6

Using the wash procedure of 10, 31, 51.5 and 71.5 mL of Example 5, the following compounds were tested for HAP substantivity: water (control); cetylpyridinium chloride (CPC); phytic acid/phosphoric acid/cetylpyridinium chloride (PaPC). The CPC was 0.0015M as described in Example 3. The PaPC solution was prepared by adding 92.6 μ L of phytic acid (40 percent weight solution, Aldrich Chemical Co., Inc.) and 20 mL of water to a 100 mL beaker. To this solution was added 3.45 g of phosphoric acid (85 percent by weight solution, Mallinckrodt) and the pH adjusted to 7.4 with the dropwise addition of 0.85 g of 50 percent by weight sodium

- hydroxide solution. The volume of the solution was then brought to 40 mL. This phytic/phosphate buffer solution was then formulated with 20 mL of 0.0045 M CPC stock solution to produce a formulation containing 0.0012 M phytic acid, 0.5 M phosphate buffer, 0.0015 M cetylpyridinium chloride. The results of the glycolysis pH test for HAP treated by the above compounds are shown in Table VI.

TABLE VI
pH as a Function of HAP Washings

mL of Water Wash	WATER	CPC ¹	PaPC ²
10	5.29	7.38	7.59
31	5.29	5.55	7.49
51.5	5.29	5.43	7.41
71.5	5.21	5.33	7.54

¹CPC = cetylpyridinium chloride

²PaCP = phytic acid, phosphate and cetylpyridinium chloride

- The results show that retention of cetylpyridinium chloride as measured by the retention of antimicrobial activity is substantially enhanced during extended water washing by the presence of phytic acid with only phosphate buffer to compatibilize the phytic acid and cetylpyridinium chloride.

Example 7

- The preparation of the HAP and the treatment of the HAP with test solutions was as previously described except in addition to using a 10 mL wash, the number of washing steps was doubled. Thus samples A, B, C and D are thus removed after washing with 22, 63.5, 84.5 and 105 mL respectively. Using this procedure, the following compounds were tested for HAP substantivity: water (control); cetylpyridinium chloride (CPC); and the following metals with phytic acid/phosphate/-cetylpyridinium chloride: calcium (Ca-PaPC); magnesium (Mg-PaPC); tin (Sn-PaPC); zinc (Zn-PaPC); strontium (Sr-PaPC); and copper (Cu-PaPC). The CPC was prepared as in Example 3. The formulations containing the metals were prepared by weighing into labeled jars the amount of metal chloride as given in Table VII

TABLE VII
Preparation of Metal Chloride Solutions

Metal	Formula	Wt. (g) added to jar	FW*
Calcium	$\text{CaCl}_2 \cdot \text{H}_2\text{O}$	0.0298	147.02
Zinc	ZnCl_2	0.0277	136.28
Strontium	$\text{SrCl}_2 \cdot 6\text{H}_2\text{O}$	0.0540	266.62
Tin	$\text{Sn(II)Cl}_2 \cdot 2\text{H}_2\text{O}$	0.0457	225.63
Magnesium	$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$	0.0411	203.31
Copper	CuCl_2	0.0273	134.45

*FW = formula weight

A 15 mL aliquot of a 0.0045 M phytic acid solution and a 15 mL aliquot of a 1.5 M phosphoric acid solution were then added to each jar and the pH adjusted to about 7 with the dropwise addition of 50 percent sodium hydroxide. A 15 mL aliquot of a 0.0045 M CPC solution was then added to each of the above solutions to produce formulations containing 0.0015 M metal; 0.0015 M phytic acid; 0.5 M phosphate; and 0.0015 M CPC. The solutions containing zinc, strontium or copper were initially water clear, with a precipitate observed after several hours.

The HAP substantivity of the formulations containing the metal as measured by the glycolysis pH test are given in Table VIII.

TABLE VIII
pH as a Function of HAP Washings¹

Volume of Wash (mL)	Water	CPC	Ca-PaPC	Mg-PaPC	Sn-PaPC	Zn-PaPC	Sr-PaPC	Cu-PaPC
22	5.12	5.92	7.59	7.56	7.55	7.61	7.62	7.61
63.5	5.22	5.24	7.67	7.66	7.64	7.71	7.72	7.66
84.5	5.33	5.34	7.75	7.77	7.69	7.78	7.81	7.73
105	5.43	5.44	7.76	7.80	7.73	7.82	7.84	7.76

¹CPC = cetylpyridinium chloride; PaPC = phytic acid, phosphate and cetylpyridinium chloride; Ca = calcium; Mg = magnesium; Sn = Tin; Zn = zinc; Sr = strontium; Cu = copper(II)

These results show that the presence of metal ions does not interfere with the ability of phytic acid to enhance the substantivity of CPC to hydroxyapatite.

Example 8

The following formulations were tested to determine the substantivity of cetylpyridinium chloride to HAP that had been treated with the test compounds in a separate step, prior to exposing the HAP to the cetylpyridinium chloride: water (control); phytic acid; sodium phosphate (NaH_2PO_4); and tin/phytic acid/phosphate (Sn-PaP). The phytic acid was used at 0.0015 M and prepared from a 0.045 M stock solution by removing a 5 mL aliquot to a vial and diluting with 10 mL of water. The pH of this solution was 2.5. The tin/phytic acid solution in 0.5 M phosphate buffer was prepared by weighing 0.0051 ± 0.0001 g SnCl_2 into a vial and brought into solution by the addition of 5 mL of water. A 5 mL aliquot of 0.0045 M phytic acid and a 5 mL aliquot of 1.5 M phosphate buffer were added to the tin solution to produce a formulation with 0.0015 M Sn; 0.0015 M phytic acid; and 0.5 M phosphate.

The HAP was prepared as described under general experimental and was treated with 2 mL of the test solutions containing phytic acid as indicated above. The HAP solids were then washed with 2-three mL portions of water and then exposed to 2 mL of 0.0015 M CPC solution. The HAP suspensions were then washed using the extended wash procedure of 22, 63.5, 84.5 and 105 mL described in Example 7.

The substantivity of the formulations as measured by the glycolysis pH test are given in Table IX.

TABLE IX
pH as a Function of HAP Washings

Volume of Wash (mL)	Water	Phytic Acid	NaH_2PO_4	Sn-PaP*
22	5.39	7.44	7.28	7.21
63.5	5.24	7.81	5.19	7.39
84.5	5.28	7.52	5.26	7.51
105	5.32	7.47	5.25	7.41

*Sn-PaP = tin, phytic acid and phosphate

The results show that the addition of tin does not interfere with the beneficial effect of phytic acid even with extended washing of the HAP. These results also show the beneficial effects of phytic acid and tin/phytic acid can be exerted by a two step exposure of the HAP.

Example 9

In this trial, the preparation of HAP and the treatment of HAP with the test solutions were as previously described under general experimental. The following solutions were tested for HAP substantivity as measured by the glycolysis pH test:

- 5 A. Water (control); cetylpyridinium chloride (CPC);
 B. Copper/phytic acid/sodium bicarbonate (CuPaB);
 C. Cetylpyridinium chloride/sodium bicarbonate (CPCB);
 D. N-tetradecyl-4-ethylpyridinium bromide (TDEP);
 E. Copper/phytic acid/sodium bicarbonate/cetylpyridinium chloride
10 (Cu-PaBC);
 F. Phytic acid/sodium bicarbonate/cetylpyridinium chloride (PaBC);
 G. Copper/phytic acid/sodium bicarbonate/N-tetradecyl-4-ethylpyridinium
 bromide
 (Cu-PaBT).

15 The concentration of the components being 0.0015 M except sodium bicarbonate at 0.5 M.

A 0.0045 M solution of N-tetradecyl-4-ethylpyridinium bromide was prepared by adding 0.0173 ± 0.0001 g of N-tetradecyl-4-ethylpyridinium bromide to a 10 mL volumetric flask and diluting to mark with water.

20 The results from this trial are given in Table X.

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TABLE X
pH as a Function of HAP Washings

VOLUME OF WASH (mL)	Samples						
	Water	CPC ¹	CuPaB ²	CB ³	TDEP ⁴	Cu- PaBC ⁵	Cu- PaBT ⁷
3	5.23	7.71	8.01	8.78	7.82	8.62	8.73
13	5.27	7.80	6.05	8.53	5.42	8.36	8.39
21	5.30	6.73	5.82	5.73	5.32	8.23	8.26
31	5.34	5.43	5.84	6.05	5.39	8.19	8.14

¹CPC = cetylpyridinium chloride;

²CuPaB = copper, phytic acid and sodium bicarbonate;

³CB = cetylpyridinium chloride and sodium bicarbonate;

⁴TDEP = N-tetradecyl-4-ethylpyridinium bromide;

⁵Cu-PaBC = copper, phytic acid, sodium bicarbonate and cetylpyridinium chloride;

⁶PaBC = phytic acid, sodium bicarbonate, and cetylpyridinium chloride;

⁷Cu-PaBT = copper, phytic acid, sodium bicarbonate and N-tetradecyl-4-ethylpyridinium bromide

The results show that sodium bicarbonate works as effectively in combination with phytic acid and cetylpyridinium chloride as does the phosphate buffer. In addition ethyltetradecylpyridinium bromide works as well as cetylpyridinium chloride as an antimicrobial in combination with phytic acid, copper and bicarbonate buffer.

5 Example 10

An *in-vivo* study was conducted to determine the ability of the test formulations containing phytic acid to inhibit the experimental formation of gingivitis in Beagle dogs.

Purebred female Beagle dogs, 2-3 years old, with naturally occurring gingivitis were randomly divided into groups of four animals each. After 14 days of adaptation, the
10 teeth of the dogs were scaled to remove supragingival calculus and polished. One week following the prophylaxes, during which time oral care was maintained by brushing, a baseline gingivitis index was obtained as measured by the procedure of Loe, *J. Periodontol.*, 38, 610 (1967) and Loe and Silness, *Acta Odont Scand*, 21, 533 (1963). After the initial gingivitis index reading, the teeth of each group were sprayed twice daily, five days per week, with
15 approximately 10 mL of one of the following mouth rinses:

- (A) cetylpyridinium chloride (CPC);
- (B) zinc/phytic acid/cetylpyridinium chloride/phosphate;
- (C) copper/phytic acid/cetylpyridinium chloride/bicarbonate;
- (D) tin/phytic acid/cetylpyridinium chloride/phosphate;
- 20 (E) phytic acid/cetylpyridinium chloride/phosphate.

All components were present at a concentration of 0.0015 M except the phosphate and bicarbonate at 0.5 M and the metals which were at 0.0014 M.

After four weeks of treatment, the gingival index was again measured. The results given in Table XI showing the increase in the gingival index, show that the phytic
25 acid/cetylpyridinium chloride/phosphate composition was the most effective composition of those tested for inhibiting deterioration of the gingival health. The phytic acid containing test solutions gave a lower increase in the gingival index when compared to the cetylpyridinium chloride alone.

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TABLE XI

Formulation ^a	Change in Gingival Index over 4 weeks
CPC	0.528
Zn-phytic/CPC/phos	0.479
Cu-phytic/CPC/bicarb	0.430
Sn-phytic/CPC/phos	0.349
phytic/CPC/phos	0.199

^aCPC = cetylpyridinium chloride; phytic = phytic acid; phos = phosphate; bicarb = bicarbonate

Example 11 Retention of CPC on Hydroxyapatite

- To separate tubes containing (1) 4.5 mM cetylpyridinium chloride (CPC); (2) 4.5 mM phytic acid, 4.5 mM CPC, and 0.5 M phosphate, pH 7.47; and (3) 4.5 mM CPC in 0.5 M phosphate was added 100 mg of washed hydroxyapatite (added as 1 mL of 100-mg/mL suspension in water). The suspensions were mixed for 10 minutes, centrifuged, a 200 μ L aliquot of the supernatant removed and diluted with 3.0 mL of water, and the absorbance of the diluted aliquot measured at 260 nanometers.
- The supernatant remaining in each tube was carefully removed and the remaining hydroxyapatite resuspended in 3.0 mL of water and mixed for 10 minutes. The suspensions were again centrifuged and a 200 μ L aliquot of the supernatant removed, diluted and ultraviolet absorbance measured at 260 nanometers. The supernatant was discarded after the absorbance reading. The procedure of resuspending the hydroxyapatite in 3 mL of water, mixing, centrifuging and measuring the absorbance of the supernatant was repeated an additional six times. By knowing how much CPC was adsorbed onto the hydroxyapatite and how much was being removed in each wash, the number of washes necessary to completely remove the CPC from the hydroxyapatite was estimated. The results from this trial are given in Table XII.

TABLE XII

Sample	Percent of *CPC initially adsorbed onto hydroxyapatite	Number of water washes calculated to remove all CPC from hydroxyapatite
CPC	37	14
Phytic acid/CPC	75	24
Phosphate/ CPC	25	14

*CPC = cetylpyridinium chloride

The results show that a compound of the present invention containing C-O-P bonds, such as phytic acid, increases the amount of CPC which is adsorbed onto the hydroxyapatite and causes the CPC to be desorbed at a lower rate.

Example 12 Compatibility of Calcium with Phytic Acid
and CPC in the Presence of Phosphate

To determine the compatibility of metal ions with phytic acid and cetylpyridinium chloride (CPC) in the presence of a phosphate buffer, a 0.0012 M solution of phytic acid was prepared in 0.5 M phosphate buffer (pH 7.32) containing 0.0015 M CPC. To separate 100 μ L aliquots of this solution was added 15, 30, 45, 60 or 75 μ L of a 0.01 M calcium chloride solution. This gave varying ratios of calcium to phytic acid in the presence of CPC in phosphate buffer. The results are given in Table XIII and show the compatibilizing effect of phosphate buffer on CPC/phytic acid solutions when exposed to calcium ions.

TABLE XIII

μ L of 0.01 M CaCl_2	Molar ratio of Phytic acid:Ca:CPC*	Observation
15	1:1.25:1.25	clear
30	1:2.5:1.25	clear
45	1:3.75:1.25	clear
60	1:5:1.25	slightly turbid
75	1:6.25:1.25	turbid

*CPC = cetylpyridinium chloride

Example 13 Effect of Surfactants on Maintaining CPC and Phytate Acid in Solution

Various amounts of the nonionic surfactants Polysorbate 80 (ICI Americas Inc., Wilmington, DE) and Poloxamer 407 (BASF Wyandotte Corp., Parsippany, NJ) and the anionic surfactant Hamposyl L-30 (WR Grace & Co., CT) were dissolved in 3.0 mM CPC solution.

Polysorbate 80 is a polyethylene oxide sorbitan ester having a molecular formula of $C_{64}H_{125}O_{26}$, an approximate molecular weight of 1309 daltons, and is available as Tween™ 80. Poloxamer 407 is a poly(oxyethylene), poly(oxypropylene) block polymer commercially available as Pluronic™ F127 from BASF and has a molecular weight of approximately 12,600 daltons.

Hamposyl L-30 is a surfactant available from W. R. Grace Company and contains by weight approximately 30 percent sodium lauryl sarcosinate, 68 percent water and 1 to 2 percent sodium laurate. Equal volumes of each of the surfactant/CPC solutions and 3.0 mM sodium phytate (pH = 7) were then mixed in small vials. The weight percent of the surfactants present in the samples ranged from zero (control) to 0.4 percent. The vials were capped and shaken and visually monitored for about 30 minutes. The contents of each vial was then transferred to a separate plastic tube and centrifuged. The supernatant was then filtered through a 0.2 micron nylon syringe filter. A 200 μ L portion of the filtrate was dissolved in 3.0 mL of water and the UV absorbance determined at 260 nm and compared against a similar aliquot of 1.5 mM CPC. The results for Polysorbate 80 indicated the lowest level of surfactant added to maintain solubility of the CPC and phytic acid under these conditions was about 0.1 weight percent; for Poloxamer 407 and Hamposyl L-30, the lowest level was about 0.2 weight percent.

Example 14 Substantivity of Phytate/CPC Solutions Compatibilized by Surfactants

The preparation of the HAP and the treatment of the HAP with test solutions was as previously described in the general experimental except the amount of HAP used was reduced to 333 μ L of 100 mg/mL suspension and the volume of water washes increased to give treated HAP samples washed with 22, 63.5, 84.5 and 105 mL of water. The following formulations from Example 13 were tested to determine the substantivity of CPC to HAP using this test:

- A. 1.5 mM CPC
- B. The filtered 0.1 weight percent Polysorbate 80.
- C. The filtered 0.2 weight percent Polysorbate 80.
- D. The filtered 0.3 weight percent Polysorbate 80.
- E. The filtered 0.4 weight percent Polysorbate 80.
- F. The filtered 0.5 weight percent Polysorbate 80.
- G. The filtered 0.2 weight percent Hamposyl L-30.
- H. The filtered 0.3 weight percent Hamposyl L-30.
- I. The filtered 0.2 weight percent Poloxamer 407.
- J. The filtered 0.3 weight percent Poloxamer 407.

The pH after incubation as a function of HAP washings for these solutions in this test are shown in Table XIV.

TABLE XIV
pH as a Function of HAP Washings

Volume of Water Wash	A*	B	C	D	E	F	G	H	I	J
22	4.92	7.22	7.15	7.12	7.02	7.09	4.82	4.36	7.22	7.16
63.5	4.86	7.21	7.17	7.14	7.14	7.12	4.54	4.49	7.27	7.21
84.5	4.93	7.21	7.17	7.17	7.14	7.14	5.37	5.43	7.25	7.22
105	4.94	7.23	7.19	7.18	7.19	7.17	4.89	5.19	7.25	7.24

*A through J correspond to the solutions having the same letter designation as given in Example 14.

Additionally, 50 μ L aliquots of the HAP samples were taken at the various stages of water washing, dissolved in 3.0 mL of 3 N HCl and analyzed by UV for CPC content by the degree of absorption at 260 nm. The results from this analysis are shown in Table XV.

- TABLE XV, -

UV Absorbance or Dissolved HAP from
Washing Experiment Which Quantifies the
Amount of CPC Still Present on the HAP

5			UV Absorbance		(ave of 2)
		22 mL wash	63.5 mL wash	84.5 mL wash	105 mL wash
	A*	0.0036	0.0031	0	0.0011
10	B	0.0666	0	0.0533	0.0380
	C	0.0622	0.0462	0.0488	0.0378
	D	0.0556	0.0533	0.0488	0.0362
	E	0.0533	0.0467	0.0436	0.0338
15	F	0.0489	0.0444	0.04	0.0331
	G	0.0084	0.0044	0.0018	0.0011
	H	0.0049	0.0031	0.0089	0.0011
	I	0.0778	0.0489	0.04	0.0216
	J	0.0689	0.0556	0.048	0.0249

*A through J correspond to the solutions having
the same letter designation as given in
Example 14.

These results show that Polysorbate 80 (Samples B through F) and Poloxamer 407
(Samples I and J) can both maintain the CPC and phytic acid in solution and do not substantially
interfere with the enhanced substantivity attributable to the presence of phytic acid. On the
other hand, Hamposyl L-30 (Samples G and H) is capable of keeping the phytic acid and CPC
mixture in solution, but interferes in the phytic acid mediated enhanced substantivity.

Example 15 Use of Sodium Bicarbonate to Maintain Phytate and TDEP in Solution

Aliquots of a 1.5 M solution of sodium bicarbonate were placed in dram vials and
diluted with water to 1 mL total volume. A 1.00 mL aliquot of 4.5 mM
N-tetradecyl-4-ethylpyridinium bromide (TDEP) aqueous solution and a 1.00 mL aliquot of a
4.5 mM sodium phytate solution were added to each vial to give 1.5 mM sodium phytate,
1.5 mM TDEP and various levels of sodium bicarbonate (between zero and 2.5 weight percent).
The caps were replaced, the vials shaken and then visually observed after standing at room
temperature (16°C to 25°C) for 30 minutes. The observed turbidity is given in Table XVI and
indicates that about 1.64 weight percent of sodium bicarbonate is needed to maintain the
sodium phytate and TDEP in solution.

TABLE XVI
Observations of Interaction Between
Sodium Phytate and TDEP in Various
Levels of Sodium Bicarbonate

Weight Percent NaHCO ₃	Visual Observation
0	Very Turbid
0.82	Turbid
1.07	Slightly Turbid
1.31	Less Turbid
1.64	Clear
2.46	Clear

Example 16 Use of Polysorbate to Maintain Phytate and TDEP in Solution

Aliquots of a 1.5 weight percent of Polysorbate 80 were placed in dram vials and diluted to 1.00 mL total volume with water. A 1.00 mL aliquot of 4.5 mM of N-tetradecyl-4-ethylpyridinium bromide (TDEP) aqueous solution and 1.00 mL aliquot of a 4.5 mM sodium phytate solution were added to the vials containing Polysorbate to give 1.5 mM sodium phytate, 1.5 mM TDEP in various levels of Polysorbate 80 (between zero and 0.5 weight percent). The caps were replaced, the vials shaken and the solutions visually monitored over time. The contents of each vial were then transferred to a plastic tube, centrifuged and the supernatant filtered through a 0.2 μ m nylon syringe filter. A 200 μ L portion of the filtrate was dissolved in 3.00 mL of water and the UV absorbance determined at 260 nm and compared to a similar sample of 1.5 mM TDEP. The results from the visual observations and the the percent of original TDEP remaining in solution after filtering are given in Table XVII.

TABLE XVII
Observations and UV Absorption Data From
the Interaction of Sodium Phytate and
TDEP in Polysorbate 80

Weight Percent Polysorbate 80	Visual Observation	Percent of Original TDEP in Solution after Filtering
0	Turbid	0
0.05	Slightly turbid	86
0.075	Clear	100
0.10	Clear	96
0.20	Clear	102
0.30	Clear	99
0.40	Clear	98

The results indicate that about 0.075 weight percent of Polysorbate maintained the phytate and TDEP in solution as visually observed and as indicated by UV absorbance.

Example 17 Use of Hamposyl to Maintain Phytate and TDEP in Solution

Aliquots of a 1.5 weight percent solution of Hamposyl L-30 were placed in dram vials and diluted with water to 1 mL total volume. A 1.00 mL aliquot of 4.5 mM of N-tetradecyl-4-ethylpyridinium bromide (TDEP) aqueous solution and a 1.00 mL aliquot of 4.5 mM sodium phytate solution were added to give 1.5 mM sodium phytate and 1.5 mM TDEP in various levels of Hamposyl L-30 (between zero and 0.5 weight percent). The caps were replaced, the vials shaken, and the solutions visually monitored over time (for approximately 1 hour). The contents of each vial were transferred to a separate plastic tube, centrifuged and the supernatant filtered through a 0.2 μ m nylon syringe filter. A 200 μ L portion of the filtrate was dissolved in 3.00 mL of water and the UV absorbance determined at 260 nm and compared to a similar sample of 1.5 mM TDEP. The results from the visual observations and percent of original TDEP remaining in solution are given in Table XVIII

TABLE XVIII
Observations and UV Absorption Data From
the Interaction of Sodium Phytate and
TDEP in Hamposyl L-30

5	Weight Percent Hamposyl L-30	Visual Observation	Percent of Original TDEP in Solution after Filtering
	0	Turbid	0
10	0.10	Slightly Turbid	82
	0.20	Clear	98
	0.30	Clear	100
	0.40	Clear	101
15	0.50	Clear	99

The results show that the amount of Hamposyl L-30 required to maintain the sanguinarine in solution is above 0.1 weight percent.

Example 18 Use of Poloxamer to Maintain Phytate and TDEP in Solution

20 Aliquots of a 1.5 weight percent solution of Poloxamer 407 were placed in dram vials and diluted with water to 1 mL total volume. A 1.00 mL aliquot of 4.5 mM of N-tetradecyl-4-ethylpyridinium bromide (TDEP) aqueous solution and a 1.00 mL aliquot of 4.5 mM sodium phytate solution were added to give 1.5 mM sodium phytate and 1.5 mM TDEP

25 in various levels of Poloxamer 407 (between zero and 0.5 weight percent). The caps were replaced, the vials shaken, and the solutions visually monitored over time (for approximately 1 hour). The contents of each vial were transferred to a separate plastic tube, centrifuged and the supernatant filtered through a 0.2 μ m nylon syringe filter. A 200 μ L portion of the filtrate

30 was dissolved in 3.00 mL of water and the UV absorbance determined at 260 nm and compared to a similar sample of 1.5 mM TDEP. The results from the visual observations and percent of original TDEP remaining in solution are given in Table XIX.

TABLE XIX
Observations and UV Absorption Data From
the Interaction of Sodium Phytate and
TDEP in Poloxamer 407

Weight Percent Poloxamer 407	Visual Observation	Percent of Original TDEP in Solution after Filtering
0	Turbid	0
0.10	Slightly Turbid	29
0.20	Slightly Turbid	64
0.30	Less Turbid	92
0.40	Clear	97
0.50	Clear	98

The results indicate that the amount of Poloxamer L-30 required to maintain the TDEP in solution is above 0.10 weight percent.

20 Example 19 Use of Polysorbate to Maintain Phytate and Sanguinarine in Solution

Aliquots of a 10 weight percent solution of Polysorbate 80 were placed in dram vials and diluted with water to 1 mL total volume. A 1.00 mL aliquot of 4.5 mM of Sanguinarine aqueous solution and 1.00 mL of 4.5 mM of sodium phytate were added to each vial to give 1.5 mM sodium phytate and 1.5 mM Sanguinarine in various levels of Polysorbate 80 (between zero and 3.3 weight percent). The caps were replaced, the vials shaken and then visually-monitored for one hour. The recorded visual observations are given in Table XX and indicate that greater than about 1.67 weight percent of Polysorbate 80 is necessary to maintain the phytate and sanguinarine in solution.

TABLE XX

Observations and UV Absorption Data
from the Interaction of Sodium Phytate
and Sanguinarine in Polysorbate 80

Weight Percent Polysorbate 80	Visual Observations
0	Turbid
0.40	Turbid
1.67	Less Turbid
3.33	Clear

Example 20 Use of Hamposyl to Maintain Phytate and Sanguinarine in Solution

Aliquots of a 3 weight percent solution of Hamposyl L-30 were placed in dram vials and diluted with water to 1 mL total volume. A 1.00 mL aliquot of 4.5 mM of Sanguinarine aqueous solution and 1.00 mL of 4.5 mM of sodium phytate were added to each vial to give 1.5 mM sodium phytate and 1.5 mM Sanguinarine in various levels of Hamposyl L-30 (between zero and 0.8 weight percent). The caps were replaced, the vials shaken and then visually-monitored for one hour. The recorded visual observations are given in Table XXI and indicate that greater than about 0.67 weight percent of Hamposyl L-30 is necessary to maintain the sanguinarine and phytate in solution.

TABLE XXI

Observations and UV Absorption Data
from the Interaction of Sodium Phytate
and Sanguinarine in Hamposyl L-30

Weight Percent Hamposyl L-30	Visual Observations
0	Turbid
0.40	Turbid
0.67	Less Turbid
0.75	Clear

Example 21 Substantivity of TDEP/Phytic Acid Solutions Compatibilized by Bicarbonate or Surfactants

The preparation of the HAP and the treatment of the HAP with test solutions was as previously described in Example 14. The following formulations were tested to determine the substantivity of TDEP to HAP using this test:

A. 1.5 mM cetylpyridinium chloride (CPC)

- 5
- B. 1.5 mM N-tetradecyl-4-ethylpyridinium bromide (TDEP)
 - C. 1.5 mM TDEP in 0.5 M sodium bicarbonate (TDEP B)
 - D. A 1.5 mM TDEP in 0.5 M sodium bicarbonate containing 1.5 mM phytic Acid (TDEPPaB)
 - E. 1.5 mM TDEP containing 0.4 percent Polysorbate 80 (TDEPPoly)
 - F. 1.5 mM TDEP containing 0.4 percent Poloxamer 407 (TDEPPolx)
 - G. Solution E but also containing 1.5 mM phytic acid (TDEPPaPoly)
 - H. 1.5 mM TDEP containing 0.4 percent Hamposyl L-30 and 1.5 mM phytic acid (TDEPPaHamp)
 - 10 I. Solution F but also containing 1.5 mM phytic acid (TDEPPaPolx)

The pH as a function of HAP washings for these solutions in this test are shown in Table XXII.

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TABLE XXII
pH as a Function of HAP Washings

Volume of wash (mL)	CPC1	TDEP2	TDEPB3	TDEP- PaB4	TDEP- Poly5	TDEP- Poly6	TDEPPa- Poly7	TDEPPa- Hamp8	TDEPPa- Poly9
22	7.09	4.79	4.98	5.09	4.97	4.97	7.24	7.18	7.11
63.5	4.67	4.84	4.79	4.79	4.99	4.99	7.25	4.81	7.12
84.5	4.78	4.85	4.89	4.78	4.98	5.01	7.25	4.82	7.12
105	4.81	4.87	4.85	4.76	4.98	5.01	7.26	4.84	4.78

1CPC = cetylpyridinium chloride

2TDEP = N-tetradecyl-4-ethylpyridinium bromide

3TDEPB = TDEP and sodium bicarbonate

4TDEPPaB = TDEP, phytic acid and sodium bicarbonate

5TDEPPoly = TDEP and polysorbate 80

6TDEPPolx = TDEP and poloxamer 407

7TDEPPaPoly = TDEP, phytic acid and polysorbate 80

8TDEPPaHamp = TDEP, phytic acid and hamposyl L30

9TDEPPaPoly = TDEP, phytic acid and poloxamer 407

Additionally, 50 μ L aliquots of the HAP samples were taken at the various water washing steps, dissolved in 3.0 mL of 3 N HCl and analyzed by UV for CPC content by the degree of absorption at 260 nm. The results from this analysis are given in Table XXIII

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Table XXIII
UV Absorbance Data from Substantivity Trial of TDEP/Phytic Acid Solutions
Compatibilized by Bicarbonate or Surfactants

Volume of Wash (mL)	CPC	TDEP2	TDEPB3	TDEP PaB4	TDEP- Poly5	TDEPPA Poly6	TDEPPA Poly7	TDEPPA- Hamp8	TDEPA- Poly9
22	.0051	.0025	.0024	.0038	.0000	.0000	.0544	.0127	.0771
63.4	.0046	.0000	.0024	.0016	.0000	.0000	.0344	.0000	.0179
84.5	.0000	.0000	.0024	.0000	.0000	.0000	.0123	.0000	.0037
10.5	.0000	.0000	.0000	.0000	.0000	.0000	.0060	.0000	.0006

1CPC = cetylpyridinium chloride

2TDEP = N-tetradecyl-4-ethylpyridinium bromide

3TDEPB = TDEP and sodium bicarbonate

4TDEPPaB = TDEP, phytic acid and sodium bicarbonate

5TDEPPoly = TDEP and Polysorbate 80

6TDEPPolx = TDEP and Poloxamer 407

7TDEPPaPoly = TDEP, phytic acid and Polysorbate 80

8TDEPPaHamp = TDEP, phytic acid and Hamposyl L30

9TDEPPaPoly = TDEP, phytic acid and Poloxamer 407

These results in conjunction with the results on the solubility of phytic acid and TDEP in the presence of bicarbonate or surfactants indicates that at certain concentrations, bicarbonate, Polysorbate 80, Hamposyl L-30, and Poloxamer 407 can all maintain TDEP and phytic acid in solution but only Poloxamer 407 and Polysorbate 80 allow the phytic acid to enhance the substantivity of TDEP to hydroxyapatite.

Example 22 Substantivity of Phytic Acid Combinations
with Sanguinarine Compatibilized by
Surfactants

The preparation of the HAP and the treatment of the HAP with test solutions was as previously described in Example 14. the following formulations were tested to determine the substantivity of Sanguinarine to HAP:

- A. 1.5 mM cetylpyridinium chloride (CPC)
- B. 1.5 mM Sanguinarine (Sa)
- C. 1.5 mM Sanguinarine containing 3.33 percent Polysorbate 80 (SaPoly)
- D. Solution C but also containing 1.5 mM phytic acid (SaPa Poly)
- E. 1.5 mM Sanguinarine containing 0.75 percent Hamposyl L-30 (SaHamp),
(The mixture was filtered immediately before use due to a precipitate being formed.)
- F. Solution E but also containing 1.5 mM phytic acid with all components being soluble (SaPaHamp).

The pH as a function of HAP washings for these solutions is given in Table XXIV.

Table XXIV
pH as a Function of HAP Washing

Volume of Wash (mL)	CPC ¹	Sa ²	SaPoly ³	SaPaPoly ⁴	SaHamp ⁵	SaPaHamp ⁶
22	7.18	5.26	5.04	6.92	7.23	7.31
63.5	4.89	5.08	5.07	4.81	5.23	4.98
84.5	4.96	5.01	5.12	4.84	5.25	4.97
105	4.98	5.11	5.09	4.89	5.31	4.98

¹CPC = cetylpyridinium chloride

²Sa = Sanguinarine

³SaPoly = Sanguinarine and Polysorbate 80

⁴SaPaPoly = Sanguinarine, phytic acid and Polysorbate 80

⁵SaHamp = Sanguinarine and Hamposyl L-30

⁶SaPa Hamp = Sanguinarine, phytic acid and Hamposyl L-30

Additionally, 50 μ L aliquots of the HAP samples were taken at the various stages of water washing, dissolved in 3.0 mL of 3 N HCl and analyzed by UV for sanguinarine content by the degree of absorption at 260 nm. The results from this analysis are given in Table XXV.

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TABLE XXV

UV Absorbance Data from Substantivity Trial of Sanguinarine/Phytic Acid
Solutions Compatibilized with Polysorbate and Hamposyl

Volume of Wash (ml)	CRC ¹	Sa ²	SaPoly ³	SaPaPoly ⁴	SaHamp ⁵	SaPaHamp ⁶
22	.0011	.0007	.0013	.0186	.0365	.1489
63.5	.0000	.0019	.0000	.0000	.0022	.0034
84.5	.0000	.0013	.0000	.0000	.0016	.0022
105	.0000	.0024	.0000	.0000	.0000	.0013

¹CPC = cetylpyridinium chloride

²Sa = Sanguinarine

³SaPoly = Sanguinarine and Polysorbate 80

⁴SaPaPoly = Sanguinarine, phytic acid and Polysorbate 80

⁵SaHamp = Sanguinarine and Hamposyl L-30

⁶SaPa Hamp = Sanguinarine, phytic acid and Hamposyl L-30

These results, in conjunction with the results on the solubility of phytic acid and sanguinarine in the presence of Polysorbate 80 and Hamposyl L-30, indicate that at certain concentrations Polysorbate 80 and Hamposyl L-30 surfactants can both compatibilize sanguinarine and phytic acid. Additionally, both Polysorbate 80 and Hamposyl L-30 allows the phytic acid to enhance the substantivity of sanguinarine to hydroxyapatite.

Example 23

A formulation containing cetylpyridinium chloride, sodium phytic and sodium bicarbonate for human trials is given in Table XXVI. This formulation was tested with water and a mouthwash which contained CPC but did not contain phytic acid. Twenty four subjects, 19 to 57 years of age were enrolled in a crossover design with 8 week duplicate measurements. At each weekly session subjects received two baseline measurements (pretreatment), then rinsed with mouthwash, and two posttreatment measurements approximately 8 hours later. Supra-gingival plaque was completely removed from preselected Ramford teeth for microbial analysis pretreatment and 8 hours post treatment. Analyses were performed on these plaque samples for aerobes, anaerobes, and fusobacteria. The ratio of the log post-treatment bacterial counts to pretreatment bacterial counts (baseline), times one hundred, is given in Table XXVII.

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TABLE XXVI
Example 23 - Formulation Used in Human Trial

5	Component	Grams	Approximate Weight Percent
	Cetylpyridinium Chloride	3.288	0.055
	Sodium Bicarbonate	252	4.2
10	Citric Acid	30	0.5
	Glycerin	300	5
	Mint Fresh Flavor	30	0.5
	Menthol Natural	1.2	0.02
15	Saccharin Powder	1.2	0.02
	Propylene Glycol	150	2.5
	Sodium Phytate (43 weight % solution)	13.6	0.23
20	Purified Water	To Make 6 Liters Volume	87

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TABLE XXVII
Percent of Remaining Bacteria 8 Hours After Rinsing

5	Bacteria Type	Treatment with Mouthwash of Example 25	Treatment with Commercially Available CPC Containing Mouthwash	Treatment with Water
	Aerobes	10.0	15.9	27.6
10	Anaerobes	12.5	21.9	27.7
	Fusobacteria	19.6	29.0	21.0

Multiple range analyses showed that the results using the mouthwash containing CPC and phytic acid was the only one significantly different from water. These results indicate that the addition of phytic acid to CPC in the presence of a compatibilizing agent (bicarbonate buffer in this formulation) can exert prolonged antibacterial effects in supragingival plaque.

Example 24 Analyses for CPC content in Plaque samples gathered in the Trial described in Example 23.

During the trial described in Example 25, plaque was harvested after 8 hours overnight from the pre-selected Ramfjord teeth. Additionally, 2 hours after treatment plaque was harvested from the opposite Ramfjord teeth. Lastly, 4 hours post-treatment samples of plaque were collected from the four second molars. All the plaque samples were placed in 1 mL of Ringers solution and later analyzed for CPC content using HPLC. For the three time points (2, 4, and 8 hour) post-treatment, the average CPC recovery for subjects using the Formulation of Example 23 was about twice that as compared to the CPC recovery for subjects using a mouthrinse containing CPC but containing no phytic acid.

Example 25 Sulfur Analysis of Breath of Human Subjects Using the Formulation of Example 23

Volatile sulfur compounds (VSC) were measured using a portable industrial sulphide monitor (model 1170, 0.5 ppm full-scale, available from Interscan Corp., Chatsworth, CA) as described by Rosenberg et al., *J. Dental Research*, 70, 1436-1440 (1991). The four day trial consisted of six volunteers that crossed over randomly from a mouthwash containing CPC (but no phytic acid) to the formulation of Example 23 (containing CPC/phytic acid/bicarbonate). The volunteers brushed their teeth at night using their usual toothpaste and toothbrush followed by rinsing for 60 seconds with 10 mL of mouthwash. The volunteers then retired for the night. The next morning the volunteers refrained from eating and drinking and minimized talking before getting their breath measured on the portable sulphide monitor

(PSM). The breath measurement was taken on the PSM after the volunteers kept their mouths closed for three minutes. Five consecutive peak values on the PSM were recorded for each volunteer for each morning reading. The morning readings were taken on consecutive days of the week (Tuesday through Friday). The average morning breath readings after use of the formula of Example 23 showed about a 50 percent reduction in volatile sulfur components as measured on the PSM when compared to average morning breathe readings after use of a mouthwash containing CPC at similar levels but containing no phytic acid.

Comparative A Use of Sodium Bicarbonate to Maintain Phytate and Sanguinarine in Solution

The procedure of Example 15 was repeated substituting a 4.5 mM sanguinarine for the TDEP solution and the amount of bicarbonate solution being between zero and 8 percent by weight. The results from the trial are given in Table XXVIII and show that even at 8 weight percent sodium bicarbonate, the phytic acid and sanguinarine were not completely soluble.

TABLE XXVIII

Observations of Interaction Between Sodium Phytate and Sanguinarine in Various Levels of Sodium Bicarbonate

Weight Percent NaHCO ₃	Visual Observation
0	Very Turbid
0.82	Turbid
1.64	Turbid
2.46	Turbid
3.28	Turbid
4.10	Turbid
8.02	Turbid

Comparative B Use of Poloxamer to Maintain Phytate and Sanguinarine in Solution

Aliquots of a 10 weight percent solution of Poloxamer 407 were placed in dram vials and diluted with water to 1 mL total volume. A 1.00 mL aliquot of 4.5 mM of Sanguinarine aqueous solution and 1.00 mL of 4.5 mM of sodium phytate were added to each vial to give 1.5 mM sodium phytate and 1.5 mM Sanguinarine in various levels of Poloxamer 407 (between zero and 3.3 weight percent). The caps were replaced, the vials shaken and then visually-monitored for one hour. The recorded visual observations are given in Table XXIX and indicate that even at 3.33 weight percent Poloxamer, the sanguinarine and phytate did not remain in solution.

TABLE XXIX

Observations and UV Absorption Data
from the Interaction of Sodium Phytate
and Sanguinarine in Poloxamer 407

Weight Percent Poloxamer 407	Visual Observations
0	Turbid
0.40	Turbid
1.67	Turbid
3.33	Turbid

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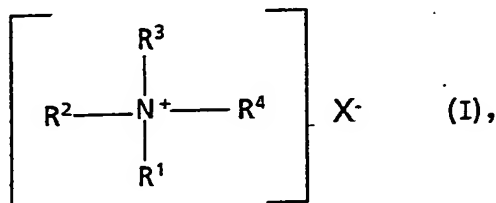
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Other embodiments of the invention will be apparent to those skilled in the art from a consideration of this specification or practice of the invention disclosed herein. It is intended that the specification and examples be considered as exemplary only, with the true scope and spirit of the invention being indicated by the following claims.

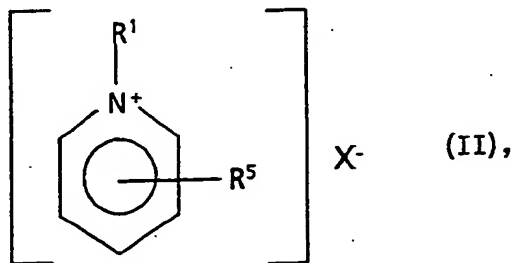
WHAT IS CLAIMED IS:

1. An oral composition comprising: (a) from 0.001 to 10 percent by weight of one or more compounds having C-O-P bonds wherein the compound having C-O-P bonds is myo-inositol hexakis(dihydrogen phosphate), myo-inositol pentakis(dihydrogen phosphate), myo-inositol tetrakis(dihydrogen phosphate) or physiologically acceptable salts thereof;
- 5 (b) from 0.001 to 10 percent by weight of one or more cationic antimicrobial compounds;
- (c) from 0.1 to 20 percent by weight of one or more compatibilizing agents; and
- (d) the remaining percent by weight is an orally acceptable vehicle.
2. The composition of Claim 1 wherein the cationic antimicrobial compound is one or more quaternary ammonium compounds of Formula I
- 10



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formula II



or a mixture thereof;

wherein

 R^1 is a C_8 - C_{20} alkyl,

R² is benzyl or C₁-C₁₂ alkyl,

R³ and R⁴ are independently a C₁-C₇ alkyl or -(CH₂-CHOH-CH₂-O)_nH wherein n is an integer from 1 to 6 inclusive,

R⁵ is -H, a C₁-C₇ alkyl or -(CH₂-CHOH-CH₂-O)_nH wherein n is an integer from 1 to 6 inclusive, and

X- is chloride, bromide, iodide or fluoride ion.

3. The composition of Claim 2 wherein the quaternary ammonium compound is cetylpyridinium chloride or N-tetradecyl-4-ethylpyridinium chloride.

4. The composition of Claim 1 wherein the cationic antimicrobial compound is sanguinarine.

5. The composition of Claim 1 wherein the compatibilizing agent is an anionic buffer, wherein the anionic buffer is phosphate, acetate, borate, citrate, bicarbonate, gluconate, tartrate, sulfate or mixtures thereof.

6. The composition of Claim 5 wherein the C-O-P compound is myo-inositol hexakis(dihydrogen phosphate) or a physiologically acceptable salt thereof; the cationic antimicrobial compound is cetylpyridinium chloride; the compatibilizing agent is bicarbonate; and the orally acceptable vehicle is 70 to 99.9 percent by weight water or an alcohol-water mixture.

7. The composition of Claims 1 wherein the compatibilizing agent is a surfactant.

8. The composition of Claim 7 wherein the surfactant is a poly(oxyethylene), poly(oxypropylene block polymer, a polyethylene oxide sorbitan ester, or a N-lauroyl sarcosine.

9. The composition according to any one of Claims 1 to 8 wherein the composition further contains a metal ion selected from copper, magnesium, tin, zinc, strontium, calcium or mixtures thereof, wherein the molar ratio of the metal ion to the compound having C-O-P bonds is from 4:1 to 1:4.

10. An oral composition comprising; (a) from 0.001 to 10 percent by weight of one or more compounds having C-O-P bonds wherein the compound having C-O-P bonds is myo-inositol hexakis(dihydrogen phosphate), myo-inositol pentakis(dihydrogen phosphate), myo-inositol tetrakis(dihydrogen phosphate) or physiologically acceptable salts thereof; (b) from 0.001 to 10 percent by weight of N-tetradecyl-4-ethylpyridinium chloride; (c) from 0.05 to 20 percent by weight of a surfactant, wherein the surfactant is a polyethylene oxide sorbitan ester; and (d) the remaining percent by weight is an orally acceptable vehicle.

11. A method of inhibiting dental calculus, dental plaque and/or oral malodor comprising administering to mammalian teeth a composition according to any one of Claims 1 to 10.

12. The method of Claim 11 wherein the mammal is a human.
13. The method of Claim 11 wherein the mammal is a dog.
14. A process for preparing an oral composition as defined in Claim 1 comprising the steps of
 - 5 (a) dissolving the compound having C-O-P bonds and the compatibilizing agent in water; and
 - (b) dissolving the cationic antimicrobial compound or a solution of the cationic antimicrobial compound in the solution obtained from step (a), wherein the pH of the solution is adjusted to between about 6 and about 8 after step (a) or step (b).
15. A kit for inhibiting the formation of dental calculus or dental plaque in a
 - 10 mammalian oral cavity comprising one or more compounds having C-O-P bonds in an orally acceptable vehicle, wherein the compound having C-O-P bonds is myo-inositol hexakis(dihydrogen phosphate), myo-inositol pentakis(dihydrogen phosphate), myo-inositol tetrakis(dihydrogen phosphate) or physiologically acceptable salts thereof; a compatibilizing agent in an orally acceptable vehicle; and one or more cationic antimicrobial compounds in an
 - 15 orally acceptable vehicle; and a means to maintain the compound having C-O-P bonds separately from the cationic antimicrobial compound.
16. The kit of Claim 15 wherein the concentration of the compound containing C-O-P bonds is from 0.001 to 10 percent by weight of the orally acceptable vehicle, the concentration of the compatibilizing agent is from 0.001 to 20 percent by weight of the orally
 - 20 acceptable vehicle and the cationic antimicrobial compound is from 0.01 to 10 percent by weight of the orally acceptable vehicle.
17. The use of a composition according to any one of Claims 1 to 10 for use as a dentifrice.
18. The use of a composition according to any one of Claims 1 to 10 for the
 - 25 manufacture of an oral composition for the treatment of dental calculus or dental plaque.
19. A commercial package comprising a compound having C-O-P bonds together with instruction for adding an antimicrobial compound and compatibilizing agent wherein the resulting dentifrice composition is as defined in any one of Claims 1 to 10.
20. A commercial package comprising an antimicrobial compound and
 - 30 compatibilizing agent together with instructions for adding a compound having C-O-P bonds wherein the resulting dentifrice composition is as defined in any one of Claims 1 to 10.

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US 92/10665

A. CLASSIFICATION OF SUBJECT MATTER

IPC5: A61K 7/16

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC5: A61K

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

FILE WPI

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	Patent Abstract of Japan, Vol 16, No 207, abstract of JP, A, 4-36229 (SUNSTAR INC), 6 February 1992 (06.02.92)	1-10, 14-20
	--	
A	EP, A1, 0329069 (COLGATE-PALMOLIVE COMPANY), 23 August 1989 (23.08.89)	1-20
	--	
A	US, A, 4335102 (S. NAKASHIMA ET AL.), 15 June 1982 (15.06.82)	1-20
	--	
A	US, A, 4305928 (K. HARVEY), 15 December 1981 (15.12.81)	1-20
	--	

☒ Further documents are listed in the continuation of Box C.

☒ See patent family annex.

* Special categories of cited documents

"A" document defining the general state of the art which is not considered to be of particular relevance

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"O" document referring to an oral disclosure, use, exhibition or other means

"P" documents published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance: the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

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"&" document member of the same patent family

Date of the actual completion of the international search

Date of mailing of the international search report

1 April 1993

21.04.93

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DAGMAR JÄRVMAN

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2

INTERNATIONAL SEARCH REPORT

International application No.
PCT/US 92/10665

C (Continuation). DOCUMENTS CONSIDERED TO BE RELEVANT		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.:
A	US, A, 4263276 (K. HARVEY), 21 April 1981 (21.04.81)	1-20

Form PCT/ISA/210 (continuation of second sheet) (July 1992)

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US 92/10665

Box I Observations where certain claims were found unsearchable (Continuation of item 1 of first sheet)

This international search report has not been established in respect of certain claims under Article 17(2)(a) for the following reasons:

1. ☒ Claims Nos.: 11-13
because they relate to subject matter not required to be searched by this Authority, namely:

See PCT Rule 39.1(iv): Methods for treatment of the human or animal body by surgery or therapy, as well as diagnostic methods.
2. ☐ Claims Nos.:
because they relate to parts of the international application that do not comply with the prescribed requirements to such an extent that no meaningful international search can be carried out, specifically:
3. ☐ Claims Nos.:
because they are dependent claims and are not drafted in accordance with the second and third sentences of Rule 6.4(a).

Box II Observations where unity of invention is lacking (Continuation of item 2 of first sheet)

This International Searching Authority found multiple inventions in this international application, as follows:

1. ☐ As all required additional search fees were timely paid by the applicant, this international search report covers all searchable claims.
2. ☐ As all searchable claims could be searched without effort justifying an additional fee, this Authority did not invite payment of any additional fee.
3. ☐ As only some of the required additional search fees were timely paid by the applicant, this international search report covers only those claims for which fees were paid, specifically claims Nos.:
4. ☐ No required additional search fees were timely paid by the applicant. Consequently, this international search report is restricted to the invention first mentioned in the claims; it is covered by claims Nos.:

Remark on Protest

- ☐ The additional search fees were accompanied by the applicant's protest.
- ☐ No protest accompanied the payment of additional search fees.

SP 8303

INTERNATIONAL SEARCH REPORT

Information on patent family members

26/02/93

International application No.

PCT/US 92/10665

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A1- 0329069	23/08/89	AU-A- 3003789	17/08/89
		DE-U- 6890139	11/06/92
		US-A- 4826675	02/05/89
US-A- 4335102	15/06/82	CA-A- 1136991	07/12/82
		FR-A,B- 2462160	13/02/81
		JP-C- 1210641	12/06/84
		JP-A- 56045407	25/04/81
		JP-B- 58046483	17/10/83
		SE-A- 8005321	01/02/81
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		AU-A- 4662179	22/11/79
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		JP-B- 63005372	03/02/88
		SE-B,C- 431395	06/02/84
		SE-B,C- 456401	03/10/88
		SE-A- 7903848	20/11/79
		SE-A- 8306210	11/11/83
		US-A- 4263276	21/04/81
US-A- 4263276	21/04/81	AU-B- 525641	18/11/82
		AU-A- 4662179	22/11/79
		BE-A- 876315	17/09/79
		CA-A- 1123339	11/05/82
		CH-A- 640728	31/01/84
		DE-A,C- 2918134	22/11/79
		FR-A,B- 2425852	14/12/79
		GB-A- 1573727	28/08/80
		JP-C- 1545238	15/02/90
		JP-A- 54154537	05/12/79
		JP-B- 63005372	03/02/88
		SE-B,C- 431395	06/02/84
		SE-B,C- 456401	03/10/88
		SE-A- 7903848	20/11/79
		SE-A- 8306210	11/11/83
		US-A- 4305928	15/12/81

Form PCT/ISA/210 (patent family annex) (July 1992)

XP-002411281

(C) WPI / Thomson

AN - 1997-281770 [26]
AP - CN19931012612 19931123
PR - CN19931012612 19931123
TI - Toothpaste used to remove dental calculus and bacterial plaque
IW - REMOVE DENTAL CALCULUS BACTERIA PLAQUE
IN - YAO H; YAO J
PA - (YAOH-I) YAO H
PN - CN1102978 A 19950531 DW199726
PD - 1995-05-31
IC - A61K7/16
DC - B04 D16 D21
AB - Toothpaste comprises special additives, an excipient and a friction agent. The special additives are a compound preparation of dismutase and epidermal growth factor and a compound preparation of zinc phytate and sodium fluoride.
- USE :
The toothpaste is used to remove dental calculus, bacterial plaque and colour stains, and for preventing dental caries, deanaphylaxis and inflammation.
- ADVANTAGE :
The pharmacodynamic stability is good and strengthens immunocompetence.

XP-002411282

(C) WPI / Thomson

AN - 1993-164329 [20]

AP - JP19910289194 19911008

PR - JP19910289194 19911008

TI - Denture cleaning agent removing dental calculus - comprises EDTA (salt) or nitrilo-tri:acetic acid, and per-carbonate salt soln. as bleaching agent

IW - DENTURE CLEAN AGENT REMOVE DENTAL CALCULUS COMPRISE EDTA SALT NITRILO TRI ACETIC ACID PER CARBONATE SOLUTION BLEACH

IN - OGAWA H; OZAWA T

PA - (LIOY) LION CORP

PN - JP5097640 A 19930420 DW199320

PD - 1993-04-20

IC - A61K7/30.

DC - D21 E16

AB - Dental calculus removing component selected from EDTA and its salt and nitrilotri acetic acid and its salt, and percarbonate salt as bleaching agent, are contained. The pH of aq. soln. contg. percarbonate at concn. 10-90 wt.% is 8-11.

- USE :

The agent has excellent dental calculus removing and cleansing effect simultaneously.

XP-002411283

(C) WPI / Thomson

AN - 2003-042450 [04]
AP - JP20010017508 20010125
PR - JP20010017508 20010125
TI - Gum for dental training for improving mastication functions in oral cavity muscle dysfunction, or corrective treatment of tooth, contains salivation accelerator and saliva buffer-capacity improver
IW - GUM DENTAL TRAINING IMPROVE MASTICATION FUNCTION ORAL CAVITY MUSCLE DYSFUNCTION CORRECT TREAT TOOTH CONTAIN SALIVA ACCELERATE BUFFER CAPACITY
IN - MIYOSHI N; TOMOTA K; TSURUI K; YUTA S
PA - (NAMI-N) NAMITECH YG
- (TOPS-N) TOP SEIKA KK
PN - JP2002220332 A 20020809 DW200304
PD - 2002-08-09
IC - A61K9/68; A23G3/30; A61C19/06; A61P1/02; A63B23/03
DC - B07 D13 D21
- P32 P36
AB - NOVELTY :
Gum for dental training contains salivation promoter and saliva buffer-capacity improver.
- USE :
For improving mastication functions in oral cavity muscle dysfunction, corrective treatment of tooth, and the restraint promotion after the corrective treatment of tooth.
- ADVANTAGE :
The dental caries preventing agent suppresses deliming of dentine and dissolution of organic substance by bacteria in the oral cavity. The gum accelerates tooth correction and provides effective mastication practice in restraint promotion after correction of tooth. The gum enables easy mastication practice in children. The hardness of gum can be changed based on the quantity of gum base used and suitable mastication recovery is enabled. The salivation promoter provides sour taste, thereby accelerates secretion of saliva, increases self-detergency in oral cavity and flushes food particles collected in the oral cavity. This provides prevention of dental plaque. The pH control by the saliva buffer-capacity improving agent enables suppression of enamel-elution. The gum for dental training provides pleasant feeling when chewing.
- PHARMACEUTICALS :
Preferred Amount: The gum for dental training contains 25-80 mass% of gum base.
Preferred Components: The gum for dental training is provided with a sugar-coating containing salivation promoter, saliva buffer-capacity improving agent and/or caries preventing agent (e.g. tannin, tea-polyphenol, fruit polyphenol, alcohol, etc).
- EXAMPLE :

A gum base (33.4 wt.%) comprising xylitol (40.1 wt.%) and maltitol (15.9 wt.%); salivation accelerator comprising citric acid (0.6 wt.%) and malic acid (0.4 wt.%); saliva buffer capacity improver comprising sodium hydrogen carbonate (2.5 wt.%), calcium hydrogen phosphate (1.8 wt.%) and tricalcium phosphate (0.5 wt.%); caries preventing agent comprising fruit polyphenol (1 wt.%); filler comprising cellulose (2.2 wt.%); and flavor comprising apple flavor (1.5 wt.%) and L-menthol (0.1 wt.%), were mixed uniformly and heated to 50 [deg]C. The heated mass was extruded in the form of thick board, then rolled to a thickness of 3 mm. The product (3 g) with the dimensions of 15 mm X 35 mm was matured for 24 hours to form a tabular gum. A sugar coating composition was prepared by compounding maltitol syrup comprising maltitol (6.5 wt.%), gum arabic (0.3 wt.%) and water (3.2 wt.%); salivation accelerator comprising citric acid (0.05 wt.%); saliva buffer capacity improver comprising sodium hydrogen carbonate (0.2 wt.%); and flavor comprising apple flavor (0.08 wt.%). The matured product was coated to 20 kg with the coating composition and gum for dental training having a weight of 1.5 g/spherical grain was obtained. 3 patients with mastication problems were treated with the gum. The gum showed improved mastication function. Saliva buffer capacity was measured and the pH was found to be 4.2 and 5.3 just before and after chewing the gum for 5 minutes. The saliva secretion amount was found to be 1 ml and 3.5 ml, before and after chewing the gum for 5 minutes, respectively. The evaluation showed that by chewing the gum salivation was promoted and pH in the oral cavity hardly fluctuated.

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